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Water Nano-Diffusion through the Nafion Fuel Cell Membrane.

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Abstract

Nafion, an amphiphilic polymer based on fluorocarbon backbones and acid groups, is probably the most widely used fuel cell membranes. According to its hydration level it self-organizes leading to nano-cavities through which water diffuse. The diffusion of water that controls the protonic transport is then central in the conversion of chemical energy to the electrical one. Recently, a sub-diffusive and heterogeneous dynamics of water were experimentally evidenced paving the way for more efficient fuel cell membranes. Fundamentally, this dynamics which occurs at the nanoscale needs to be microscopically understood. Molecular dynamics simulations are thus performed to locally examine the water dynamics and its relation with the water and Nafion structure. The sub-diffusive regime is numerically corroborated and two sub-diffusive to dif-

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diffusive transitions are found. The first is time dependent whereas the second is rather water uptake dependent. The sub-diffusive regime is ascribed to the water molecules strongly anchored close to the acid groups. We show that the sub-diffusive to diffusive time transition is the result of the elapsed time before to escape from the attractive interactions of the acid groups. The diffusive regime is recovered far from the acid groups in a homogeneous water phase that is the result of the percolation of the hydrogen bonding network. The progressive transition between sub-diffusive to diffusive regime as a function of the hydration level is due to the respective weight of diffusive dynamics that increases with respect to the sub-diffusive regime given the increase in diffusive pathways as the expense of the localized dynamics. Close to the fluorocarbon backbones the dynamics of water is also sub-diffusive but faster whereas time dynamical transition is not observed. Furthermore we highlight the existence of water corridors based on the hydrogen bonds between molecules forming single file structure in line with a sub-diffusive dynamics. These water corridors are thus possible conducting pathways of protons in a Grotthuss mechanism. Structurally we depict an interdigitated structure where the sulfonate are interleaved. Eventually, at high water uptake, we exhibit the self-organizing of Nafion leading to a phase separation between water and the Nafion membrane. We establish a specific interfacial organization of the hydrophilic sulfonate groups involving a water/Nafion interface.

Keywords: Nafion, Water diffusion, molecular simulation, anomalous dynamics

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1. Introduction

Nowadays the development of new sources of energy is one of major challenges of our society. To overcome this penury due to the intensive exploitation of our natural resource, new approaches have to be promoted in a context of sustainable development. One of the keys of this energetics problem could be

the improvement and the use of fuel cells. These are compact, light and quiet with less pollution emission. Additionally, fuel cells incorporate solid materials inducing an easy assembly and are capable to deliver a strong power with a large current density and a very good energetics performance.

10 Fuel cells are electrochemical devices allowing a direct conversion of the chemical energy provided from a fuel as hydrogen, methanol or hydrocarbons in electrical energy. As illustrated in Figure 1 fuel cells are built from two electrodes separated from an electrolytic ionic conductor. According to the operating temperature and then to the application (mobile or stationary applications,
15 high or short power supplied), the electrolyte will be different. Two classes of fuel cell can be defined i) fuel cells with an operating temperature beyond 250°C such as the solid oxide fuel cell or the Molten-carbonate fuel cell and ii) fuel cells with an operating temperature below 250°C such as Proton Exchange Membrane Fuel Cell (PEMFC) or Direct-methanol fuel cell [1].

20 PEMFC are essential given their high performance, no emission in CO₂ and an use in a large range of power. Several studies have been devoted these two past decades to reduce the cost and to increase their lifetime to allow a large commercialization. Suitable functioning needs are: i) a diffusion of gases and condensation/vaporization of water into the pores of the active layers and dif-
25 fusion layers, ii) an electronic transport and electrochemical conversion of the electrons, protons and reactive gases in the active layer of the electrodes and iii) transport of protons through the fuel cell membrane. The current membranes are based on perfluoro-sulfonic membranes and are built from fluorocarbon backbones and acidic groups. The oldest, most studied and used membrane
30 is probably the one manufactured from Dupont de Nemours (US); the Nafion membrane.

Structure of hydrated Nafion is strongly investigated in order to improve its fuel cell performance but also to design new membranes. Structural studies have highlighted a complex and multi-scale structure [2, 3, 4, 5, 6, 7, 8, 9, 10, 11, 12].
35 Additionally, a phase segregation between hydrophobic and hydrophilic moieties of Nafion was evidenced during the hydration process [3]. One of most impor-

tant properties of Nafion is its conductivity that is strongly connected to the hydration level and the water transport into the membrane. Indeed, as shown Figure 1 the water uptake progressively increases due to the reactivity between the protons and O_2 involving formation of water molecules. Therefore, an understanding of the mechanisms of conduction as a function of the hydration level through the membrane is crucial to design new performing membranes. Transport of protons through the Nafion membrane is then possible through the Grotthuss mechanism [13] and a vehicular one [14] implying chemical reactivity. Although this reactivity is crucial in both mechanisms, the role of water structure and the topology of the Nafion membrane, that are strongly connected to the hydration level, must be clarified because they control the conduction pathway allowing the protonic transport.

Due to progressive water filling three states of hydration can be defined, i) at low hydration (3 water molecules per sulfonate group) where water molecules solvate sulfonate groups of Nafion and protons leading to a localized water structure known as bounded (B) water, ii) at the intermediate hydration (20 water molecules per sulfonate group) such that water molecules are still close to H^+ and SO_3^- but less anchored and can be considered as loosely bounded (LB) and iii) at high water loading (beyond 20 water molecules per sulfonate group) where the water molecules far from H^+ and SO_3^- are then considered as free (F) [15, 16, 17]. Additionally, it has been evidenced that Nafion membrane self-assembles upon hydration leading to a peculiar topology based on channels network involving specific water conduction pathways (between channels)[6, 10, 11, 12]. An understanding of the water network and the water transport into Nafion is then essential to well apprehend the protonic conduction.

As previously evoked Nafion is a multi-scale material and the physical phenomena ruling the water transport occur from nanoscale to macro-scale [9]. Interestingly, it is now well established that mechanisms at the nanoscale strongly impact the macroscopic properties and it is then necessary to well characterize and understand the water structure and dynamics at the nanoscale. Since years molecular simulations has become an useful method to capture the micro-

scopic details and to connect them with macroscopic properties using statistical physics. Several numerical and experimental works have been performed to
70 unveil the water diffusion and the so-associated mechanism as well as the atomistic structure [18, 19, 20, 21, 22, 23, 24, 25, 26, 12]. Indeed, by using Nuclear Magnetic Resonance (NMR) pulsed gradient spin experiments, C. Casieri et al. have shown a sub-diffusive dynamics of confined water at the micrometer scale for temperatures $> 320\text{K}$ and at different water uptakes [27]. At the
75 same time, D.T. Hallinan et al. highlighted a non-Fickian diffusion of water in Nafion from time-resolved Fourier transform infrared-attenuated [21]. More recently, Mabuchi and Tokumasu have also established a sub-diffusive dynamics of protons and water at low water contents below 5 ns [28]. According to these authors, from 5 ns the diffusion became diffusive and the self-diffusivity
80 can be calculated. Devanathan et al. have also investigated the relation between structure and dynamics of water and ions. By using classical molecular dynamics simulation these authors have evidenced that at low hydration levels, coordination of hydronium ions by multiple sulfonate groups prevents vehicular transport and impedes structural transport of protons through steric hindrance
85 to hydration of the hydronium ions.[29] Devanathan et al. have also highlighted from quantum hopping molecular dynamics the existence of the water network percolation in Nafion and its effect on proton conductivity. [29].

Although these works provide insights into translational dynamics, the microscopic mechanisms and the relationship between the atomistic structure of
90 the Nafion and dynamics of water were slightly examined. However, recently Song et al. have depicted that the surface chemistry of Nafion strongly impacts the water and ions transport. Indeed, using molecular spin probes that selectively partition into heterogeneous regions of Nafion and Overhauser dynamic nuclear polarization relaxometry, Song and co-workers have revealed that
95 both water and proton diffusivity are significantly faster near the fluorocarbon and the acidic groups lining the water channels compared to within the water channels [25]. More recently, Q. Berrod et al. have also studied water motion and ion transport by combining Quasi Elastic Neutron Scattering, Pulsed Field

Gradient NMR, and atomistic simulations. They demonstrated that confinement at the nanoscale and the direct interactions with the charged interfaces produce anomalous sub-diffusion, due to a heterogeneous space-dependent dynamics within the ionic nano-channels [9].

In this work, we aim to investigate the translational diffusion of water at the nanoscale and to capture the microscopic processes ruling it. To do so molecular dynamics (MD) simulations of hydrated Nafion were carried out for different hydration states.

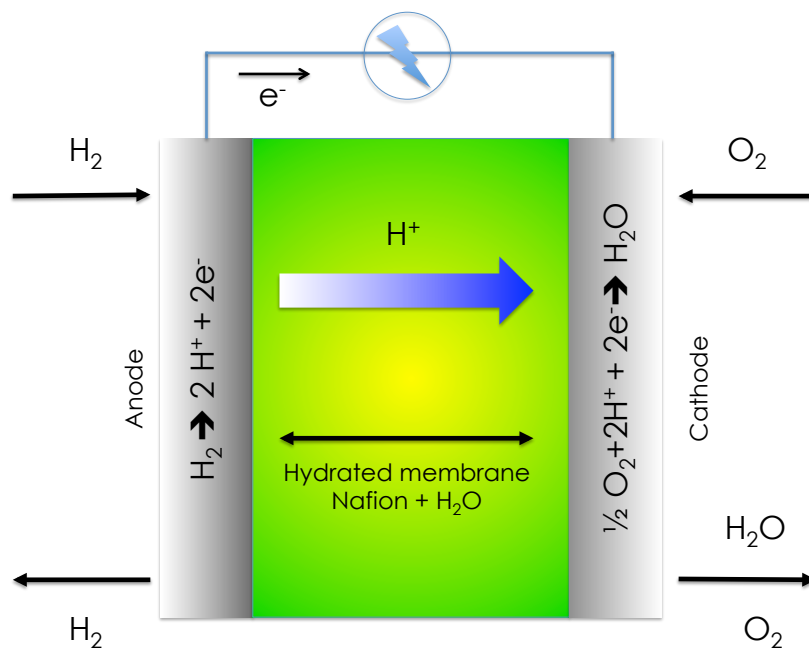


Figure 1: Illustration of the operating fuel cell membrane.

2. Methodology

2.1. Models and Computational Details

110 Nafion fuel cell membrane is a dense polymer and its molecular modeling needs specific construction to reach a relaxed configuration. To do so, Nafion membrane was built by considering 4 chains of 682 atoms of 10 monomers (see Figure 2a) by using a Monte Carlo construction based on the statistical bias developed by Theodorou & Seuter [30]. In a second time, we equilibrated the

115 empty polymer (dehydrated Nafion polymer) using the 21 MD steps scheme proposed by Hofmann et al. [31]; Within this scheme, seven cycles of three MD simulations were performed. (1) NVT, $T=600$ K; (2) NVT, $T=300$ K and(3) NpT, $T=300$ K, where N is the number of atoms, V the volume and T the temperature. Details of each steps are provided in Table 1. The value of the pressure

120 in the NpT simulations was gradually increased, passing from the first to the third cycle up to $p_{\max} = 50$ kbar, and then was reduced stepwise until it reached the final value of $p = 1$ bar. In a last step, membrane was subsequently hydrated and 7 hydration levels were defined from the number of water molecules around a sulfonate group (λ), $\lambda = 3, 5, 7, 9, 11, 14.5, 22.5$. All hydrated mem-

125 branes were relaxed from 21 MD steps scheme by using similar temperature and pressure that ones used in case of the empty polymer. Hydration was performed by randomly inserted water molecules within the relaxed empty polymer and equilibration of hydrated Nafion in statistical NpT ensemble was carried out at 1 bar and 300 K. During all processes polymer was considered as flexible.

130 Interestingly, isotherm of adsorption and water activities could be extracted from molecular dynamics simulations in the Grand Canonical statistical ensemble in order to consider the flexibility of the polymer. However, this task needs additional and specifics developments to carry out Grand Canonical Molecular Dynamics simulations such as the methodology of hybrid osmotic Monte Carlo

135 simulation[32]. That is in progress and is out of the scope of this work where the water uptake is fixed.

Nafion was modeled by considering the intramolecular and intermolecular

Table 1: Details of the thermodynamic conditions for the 21 MD steps used in the process of generation of relaxed Nafion polymer.

| Step | ensemble | T (K) | p (kbar) | Time (ps) |
|------|----------|-------|----------|-----------|
| 1 | NVT | 600 | | 50 |
| 2 | NVT | 300 | | 50 |
| 3 | NpT | 300 | 1 | 50 |
| 4 | NVT | 600 | | 50 |
| 5 | NVT | 300 | | 100 |
| 6 | NpT | 300 | 30 | 50 |
| 7 | NVT | 600 | | 50 |
| 8 | NVT | 300 | | 100 |
| 9 | NpT | 300 | 50 | 50 |
| 10 | NVT | 600 | | 50 |
| 11 | NVT | 300 | | 100 |
| 12 | NpT | 300 | 25 | 5 |
| 13 | NVT | 600 | | 5 |
| 14 | NVT | 300 | | 10 |
| 15 | NpT | 300 | 5 | 5 |
| 16 | NVT | 600 | | 5 |
| 17 | NVT | 300 | | 10 |
| 18 | NpT | 300 | 0.5 | 5 |
| 19 | NVT | 600 | | 5 |
| 20 | NVT | 300 | | 10 |
| 21 | NpT | 300 | 0.001 | 2000 |

DREIDING force field [33] that was already previously validated [34, 2]. H_3O^+ and water molecules were modeled from the model of Baaden [35, 36] and from

140 the TIP4P-2005 force field [37], respectively. The interactions between the poly-
 mer, water and proton were taken as a combination of the electrostatic and van
 der Waals (VdW) interactions. VdW contributions were modeled by consider-
 ing the Lennard-Jones potentials, such that cross-interactions were calculated
 by means of the Lorentz-Berthelot mixing rules. The electrostatic interactions
 145 were computed by considering the Ewald sum [38, 39]. All interactions were
 truncated by using a cutoff of 12 Å. All of the simulations were performed using
 the DLPOLY software [40]. Verlet-velocity propagator combined with the Nose-
 Hoover thermostat and barostat were here considered [41]. Relaxation times of
 0.5 ps of both thermostat and barostat were used. Acquisition phases were
 150 conducted during 20 ns with a timestep of 1 fs while the equilibration phases
 were lasted 50 ns. To ensure that the systems are mechanically equilibrated the
 local pressures of three components according to z direction were calculated.
 Indeed, the mechanical equilibrium is reached when the local pressures such as
 $P_x(z)$, $P_y(z)$ and $P_z(z)$ are constant as a function of z [42]. Pressure were cal-
 155 culated on 20 points by using the Irving-Kirkwood method [] and the averages
 on these 20 points with their uncertainties were calculated. We found thus the
 following values $\langle P_x(z) \rangle = 0.12 \pm 0.02$ Mpa; $\langle P_y(z) \rangle = 0.16 \pm 0.02$ Mpa and
 $\langle P_z(z) \rangle = 0.09 \pm 0.02$ Mpa in good agreement with the input pressure of 0.1
 Mpa. These small uncertainties show that the pressure is well converged, con-
 160 firming that the simulated systems have reached a mechanical equilibrated state.
 These verifications were also performed for the hydrated membranes where the
 uncertainties were found slightly higher around 0.05 Mpa that is sufficient to
 conclude on the mechanical equilibrium. Let us mention that these values of
 uncertainties are the same order of magnitude than those found in case of liquid
 165 [43, 42].

We report in Figure 2b the density of the hydrated membrane as a function of
 λ . As pointed out in Figure 2b a fair agreement with experiment was obtained
 that validated the so-used force fields of Nafion, proton and water and their
 combining.

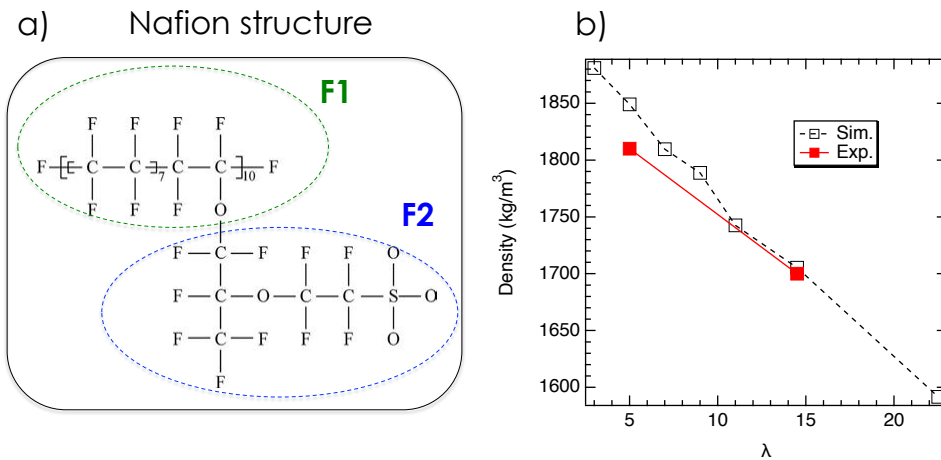


Figure 2: a) Structure of the Nafion membrane. F1 and F2 are the labels of fluorine atoms in the backbone and branching, respectively. b) Density of the hydrated membrane as a function of λ .

170 2.2. Translational dynamics

Translational dynamics was studied from the calculation of the mean square displacement (MSD)

$$\text{MSD}(t) = \frac{\left\langle \sum_{t_0} \sum_{i=1}^N [\mathbf{r}_{\text{com},i}(t+t_0) - \mathbf{r}_{\text{com},i}(t_0)]^2 \right\rangle}{NN_0 t} \quad (1)$$

with $\mathbf{r}_{\text{com},i}$ the position of the centre of mass of molecule i , t_0 the time origin, N the number of molecules and N_0 the number of t_0 . From the MSD and by
 175 using the generalized Einstein's relation, the diffusivity (D_f) is then extracted such that

$$\text{MSD}(t) = D_f t^\alpha \quad (2)$$

α is a parameter related to the type of diffusion, $\alpha = 1$ corresponds to a diffusive regime, $\alpha < 1$ is connected to a sub-diffusive one while $\alpha > 1$ is linked to a super-diffusive process. The case of $\alpha=2$ corresponds to a ballistic
 180 dynamics [44, 9, 45]. Let us mention, that for $\alpha \neq 1$, the diffusion is considered as anomalous. For a diffusive (Fickian) regime the diffusivity is related to well

known diffusion coefficient (D) extracted from the Einstein's relation in one dimension ($MSD(t) = 2Dt$). α was calculated by deriving Eq. 2 as

$$\alpha(t) = \frac{\partial MSD(t)}{\partial t} \cdot \frac{MSD(t)}{t} \quad (3)$$

3. Results and Discussion

185 We report in Figure 3 the instantaneous value of α as a function of time (t) for all water uptakes. Figure 3a shows that for $t < 1$ ns, $\alpha < 1$ that suggest a sub-diffusive regime in this time range. As highlighted in Figure 3b, beyond 1 ns, α rises to one leading to a crossover between a sub-diffusive to a diffusive regime (or a lesser sub-diffusive regime). Interestingly, this effect seems to widen
190 as the water uptake increases to strive for a quasi diffusive regime. At low λ , we can assume that water molecules are probably organizing in clusters leading to the formation of nanophases where molecules locally diffusive (because they are trapped in small size environment) inducing a sub-diffusive regime. Indeed, microscopically speaking sub-diffusive regime is the result of the local diffusion
195 and a heterogeneous structure. In second stage size of nanophases could growth as λ increases involving a progressive transition toward a diffusive regime. The water structure and its connection with the dynamics will be deeply discussed later.

We report in Figure 4a the average value of α ($\langle \alpha \rangle$) as a function of λ
200 calculated between 0 and 1 ns. As shown in Figure 4a the sub-diffusive dynamics is clearly evidenced ($\langle \alpha \rangle < 1$). Additionally, Figure 4a shows an increase in $\langle \alpha \rangle$ as the water uptake increases that highlight a continuous dynamical transition. To well appreciate the loss of the sub-diffusive regime beyond 1 ns, $\langle \alpha \rangle$ was calculated between five consecutive time intervals noted Δt_i ($\Delta t_i = 1$
205 between 0 and 1 ns; $\Delta t_i = 2$ between 1 and 5 ns; $\Delta t_i = 3$ between 5 and 10 ns; $\Delta t_i = 4$ between 10 and 15 ns and $\Delta t_i = 5$ beyond 15 ns). Figure 4b displays α as a function of Δt_i for six water contents. Figure 4b exhibits a progressive increase of $\langle \alpha \rangle$ as a function of Δt_i for all loadings that is in line with previous

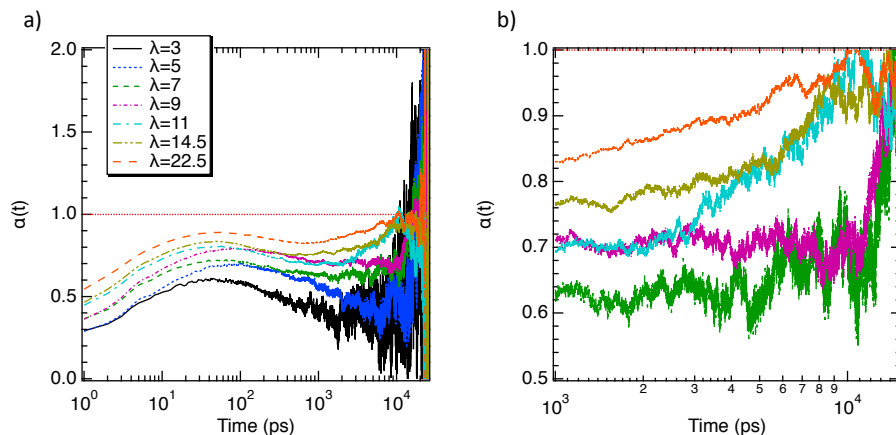


Figure 3: α as a function of time for all hydration levels in all range of time a) and between 1 ns and 15 ns b).

conclusions drawn from instantaneous values of α . This effect is found higher at
 210 low loadings because from $\Delta t_i = 3$, $\langle \alpha \rangle$ is doubled. The diffusion coefficients
 were extracted from the linear region of the MSD related to the diffusive regime
 and from the total MSD trajectories (from 0 to 20 ns) and reported in Figure
 4a. When the diffusion coefficient is evaluated from all MSD trajectories a fair
 agreement between experiment [46] and simulation was found. Indeed tendency
 215 as a function of λ and absolute values are well reproduced. Experimentally
 it was not possible to de-correlate both regimes. Interestingly, when only the
 diffusive regime is considered the so-calculated diffusion coefficients are found
 weaker suggesting that molecular simulations could be a relevant tool to be
 combining with NMR or QENS experiments to extract the diffusion coefficient
 220 of confined liquids. To sum up; i) the sub-diffusive regime could be due to the
 presence of nanophases inducing localized dynamics and ii) the loss of the sub-
 diffusive regime could be the result of to spatial connection between nanophases
 inducing the diffusive pathways. We first began then to examine the existence
 of nanophases and their structure.

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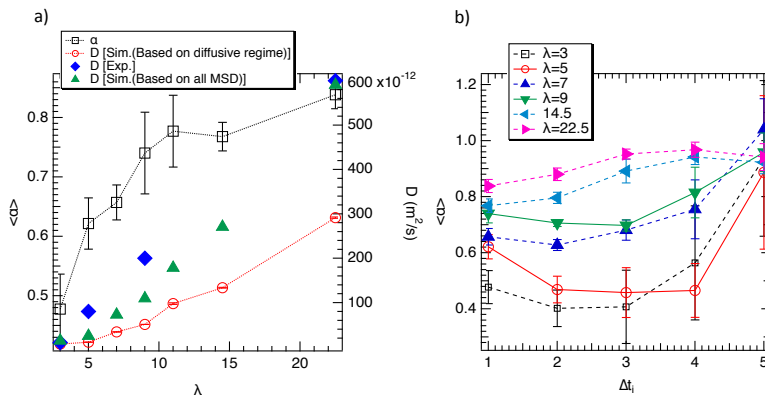


Figure 4: a) $\langle \alpha \rangle$ (left axis) and translational diffusion coefficient (right axis) as a function of λ . b) $\langle \alpha \rangle$ as a function of Δt_i . Experimental diffusion coefficient were taken in Ref. [46].

Water nanophases can be compared to the water pockets and can be assimilated to the clusters defined from their size. By using the method of Stoddard [47] the cluster size distribution was calculated by considering the hydrogen bonds (HB). HB were defined from the geometric criteria established by Luzar and Chandler [48] that consist in the distances of 2.5 Å and 3.5 Å between the hydrogen and the oxygen atoms and two oxygen atoms of two water molecules.

230 We report in Figure 5a the cluster size distribution of water as a function of λ . At low water uptake ($\lambda = 3$ and 5) all cluster sizes are equally sampled and the system does not percolate. This result sheds light on a polydisperse distribution

235 of clusters and then the presence of small water islands. The percolation occurs from $\lambda=7$ where all nanophases begin to be gathered in a single interconnected structure. This structural organization is in line with the so-observed transition between sub-diffusive to diffusive regime. Interestingly, despite the percolation of the HB network that occur at $\lambda = 7$, small clusters are also evidenced suggesting a mixing between the single interconnected structure and clusters, that

240 is in good concordance with the so-observed sub-diffusive and diffusive regime. To provide an obvious picture of time evolution of clusters, the cluster size distribution was evaluated for each Δt_i . As shown in Figure 5b, for $\lambda = 3$ the

cluster size distributions is not time-dependent that suggest an absence of cluster interconnection and an absence of clusters assembly with respect to the time. Therefore, at low λ , water molecules are trapped in the nanophases and locally diffusive in small size environment inducing a sub-diffusive regime.

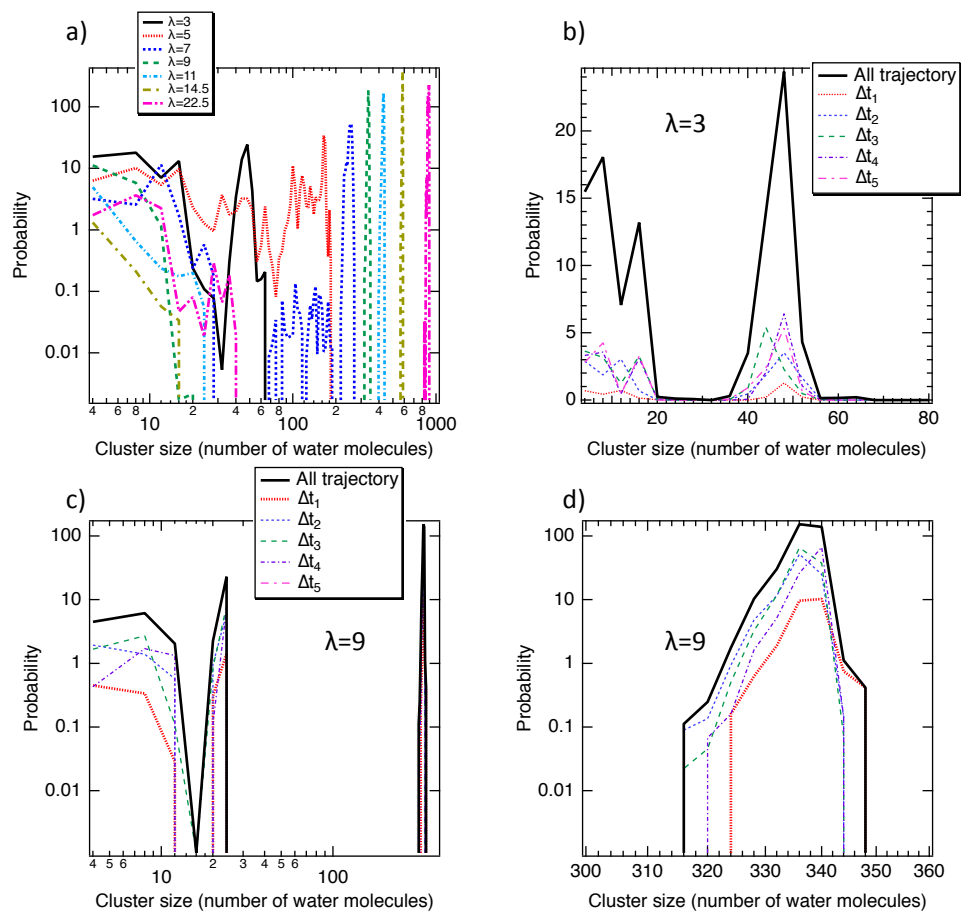


Figure 5: a) Number of cluster as a function the cluster size for all hydration levels. Number of cluster as a function the cluster size for $\lambda = 3$ b) and $\lambda = 9$ c). d) Enlargement of part c) for high size clusters.

To illustrate the water nanophases, snapshots of final configuration of six hydration levels are reported in Figure 6. As shown in Figure 6, at low water uptake water clusters are observed to be located close to the sulfonate groups

whereas at higher loadings water molecules are rather connected to form a structuring network always impacted by the hydrophilic sulfonate groups and hydrophobic fluorine backbone. It would seem that water molecules firstly adsorb on the sulfonate groups and grow to form water nanophases. As λ increases nanophases grow to finally interconnect and form a single structure. The localized dynamics are thus always present but they statistically decrease in favor to the diffusive regions (nanophases assembly). Indeed, as shown in Figure 5c and 5d where the cluster size distributions for $\lambda = 9$ is represented, the presence of small clusters and the percolation phenomenon are observed whatever Δt_i . That suggest the permanent presence of the single interconnected water structure. This result refutes the fact that the sub-diffusive to diffusive transition is due only to the time assembly of nanophases [49].

A question arises on the origin of the sub-diffusive regime: is the result of the confined dynamics into nanophases or the local diffusion close to the sulfonate groups? Another point concerns the loss of the sub-diffusive regime from $\lambda = 5$ that can be correlated to the assembly or a growth of water nanophases leading to an increase in diffusive domains with respect to the sub-diffusive ones. This scenario does not exclude the presence of a localized dynamics whose statistical weight will decrease as the water nanophases grow.

To analyze the local dynamics close to the sulfonate groups (SO_3^-) the characteristic distances between water and SO_3^- were examined. To do so, the radial distribution functions (RDF) between hydrogen atoms of water molecules and oxygen atoms of sulfonate groups are calculated and reported in Figure 7a. As shown in this figure two maxima can be extracted whatever the water uptake, one around 2.0 Å corresponding to a strong hydrogen bond and a second distance located at 3.4 Å connected to the second hydration shell. Difference in intensity between the three water loadings is due to a density effect [50]. RDF between oxygen atoms and fluorine atoms were also calculated to investigate the dynamics along the hydrophobic backbone. Contrary to the sulfonate groups, the first maximum is found at 4 Å for both types of fluorine (noted F1 and F2

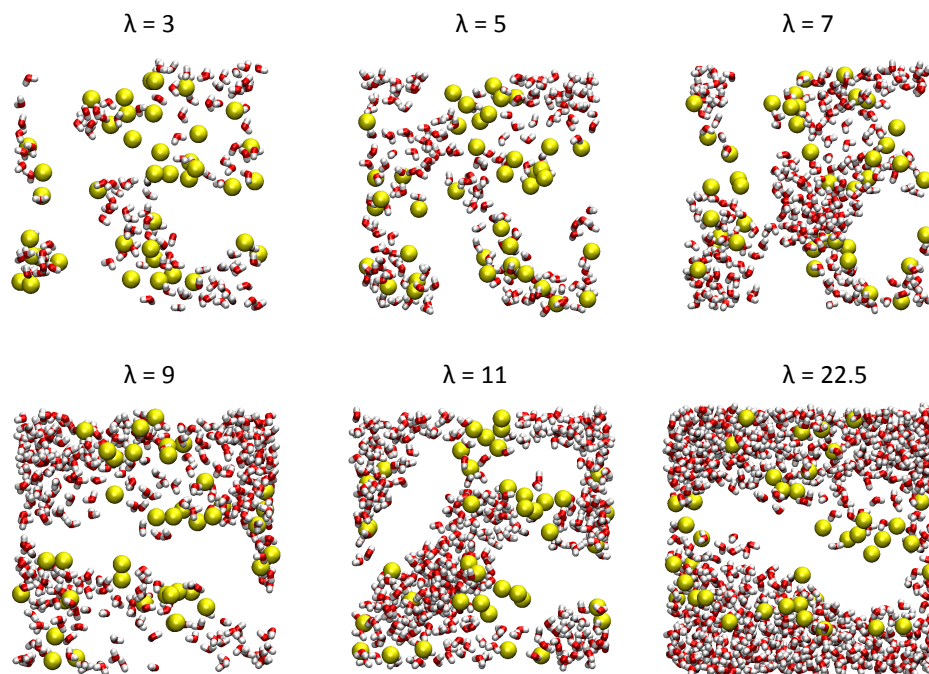


Figure 6: Snapshots of final configurations of 6 hydration levels. water molecules are represented such that red and white colors represent the oxygen and the hydrogen atoms, respectively. Yellow spheres are the sulfur atoms in the Nafion membrane. For clarity, other atoms of Nafion are not represented.

in Figure 2) whereas the position of the second peak changes as a function of type of fluorine. This difference is due to the proximity of the sulfonate groups. Indeed as illustrated in Figure 2, F2 is close to the sulfonate groups and then
 285 benefit from their hydrophilicity. From RDF, three regions around the sulfonate groups were drawn, $< 2.0 \text{ \AA}$ corresponding to the bounded water, between 2.0 \AA and 3.4 \AA corresponding to the loosely bounded water and beyond 3.4 \AA that is the free water. Similar decomposition was performed around the fluorine. Local dynamics have been then investigated by considering this spatial division.

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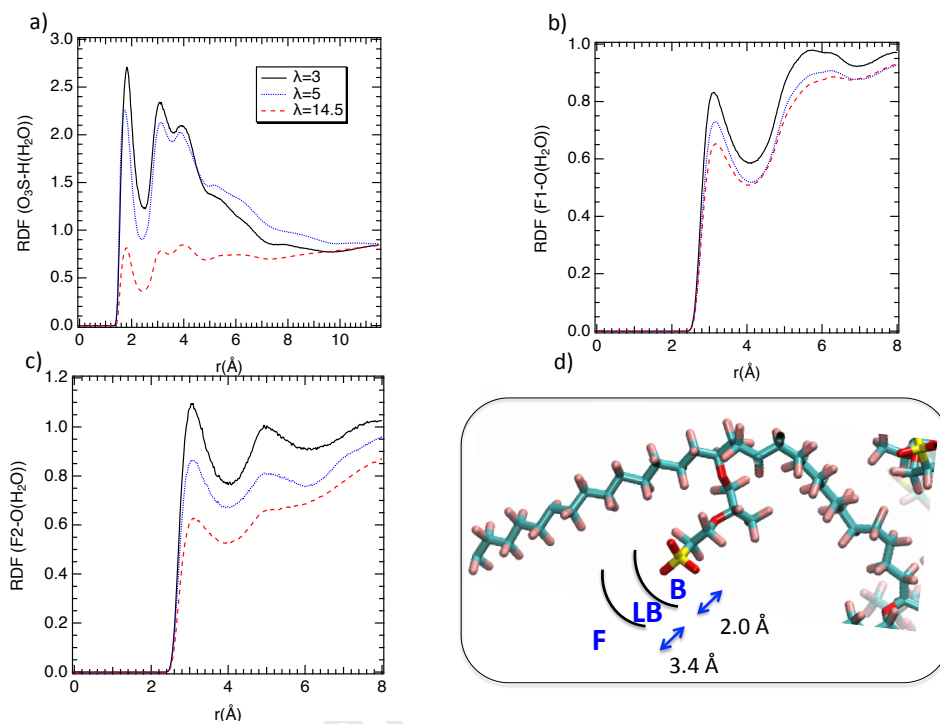


Figure 7: a) Radial distribution functions between oxygen atoms of SO_3^- groups of the Nafion and the hydrogen atoms of water for three water uptakes. RDF between F1 atoms b) and F2 atoms c) (labels are defined in Figure 2) and oxygen atoms of water. d) Illustration of the so-considered distances, based on the RDF calculation, to define three regions around the sulfonate groups; bounded water (B), loosely bounded water (LB), and free water (F).

We report in Figure 8 the MSD of four water uptakes from $\lambda = 3$ to $\lambda = 22.5$ in three bounded (B), loosely bounded (LB) and free (F) water regions. As highlighted in Figure 8a a sub-diffusive regime is observed in the three zones because all water molecules are anchored close to the sulfonate groups (see Figure 6 and Figure 7a). In both LB and F boundaries the sub-diffusive regime is progressively replaced by the diffusive one as the water uptake increases. That is the result of the assembly of nanophases allowing a diffusive regime through the membrane (see Figure 6). That is an evidence of the connection between the percolation process (Figures 5 and 6) previously highlighted and the water diffusion. Indeed, we have shown that the diffusion of water molecules increases

with the water uptake i.e. with the rate of percolation and the growth of the hydrophilic part in line with the work of Benziger et al. who showed that the different percolation thresholds suggest the hydrophilic domains in Nafion grow from lamella[51]. However as depicted in Figure 8c and 8d the dynamics of water in the B region is still sub-diffusive during about 50-100 ps prior to becoming diffusive. The macroscopic sub-diffusive regime previously characterized by α could be then attributed to the water molecules strongly anchored close to the sulfonate groups. The time elapsed in the sub-diffusive regime could be probably attributed to the elapsed time before to escape from the attractive interactions of the sulfonate groups. Indeed, during the first hundred pico-seconds the MSD was found around 2-5 Å² that represent a displacement of 2.2 Å that is the same order of magnitude of size of region B.

As shown in Figure 9a this phenomenon is independent of the water uptake that bear out an escaping process of water molecule close to the sulfonate groups. The elapsed time of water molecules close to the sulfonate groups can be quantified by calculating their residence times from the evaluation of the correlation function of the survival probability function ($P(t)$) that adopt a value of one if the water molecule labeled j has been in the referred hydration shell around site α from time t' to time $t + t'$, without getting out in between this time interval, and zero otherwise. This function is then fitted from a double exponential such that the residence time is extracted [52, 53]. We report in Figure 9b the correlation function as a function of time for all λ . As shown in Figure 9b the decorrelation is obtained for $t \sim 50$ ps that is in line with the predicted time transition between both dynamical regimes. Residence times were calculated from an exponential adjustment using a double exponential. As reported in Figure 9b the fit is in good concordance with $P(t)$ that suggest that the residence time can be extracted. We obtain the following residence times, $\tau_{\lambda=22.5}=84.4$ ps and $\tau_{\lambda=3}=63.7$ ps that is line with the time crossover observed in Figure 8.

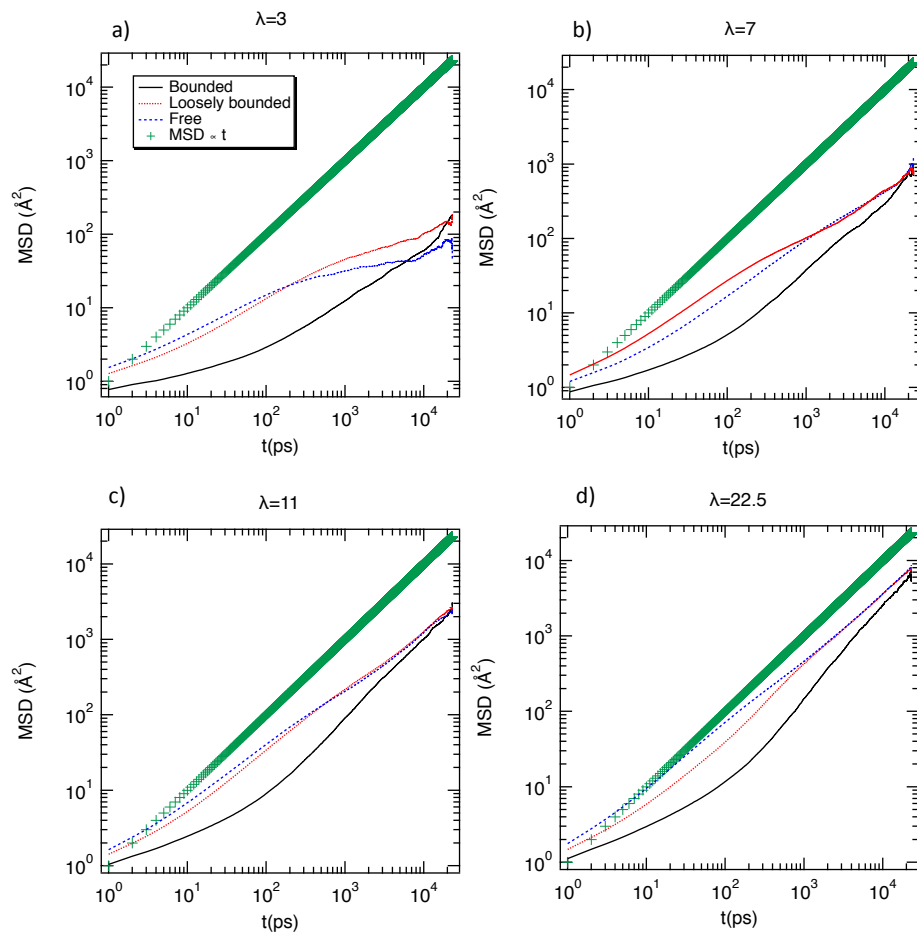


Figure 8: Mean square displacements of water in three water regions B, LB and F around SO_3^- for a) $\lambda = 3$, b) $\lambda = 7$, c) $\lambda = 11$ and d) $\lambda = 22.5$.

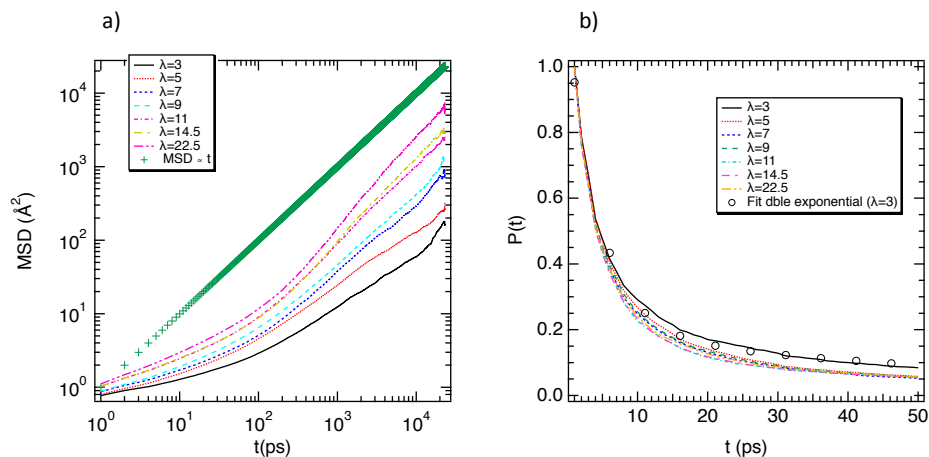


Figure 9: a) Mean square displacements of water in the region B around SO_3^- for all uptakes. b) Correlation function of the survival probability function as a function of time allowing us the calculation of residence times.

330 Interestingly, as shown in Figure 10, the dynamics of water around fluorine atoms is found similar in three regions due to the hydrophobicity of carbon-fluorine backbone preventing the anchoring of water molecules. As for the water molecules close to the sulfonate groups a progressive transition between sub-diffusive to diffusive dynamics as the water uptake increases is also evidenced.

335 Indeed, the minimum distance between the fluorine atoms and the sulfonate groups is found between 3.2 \AA and 3.9 \AA (from RDF calculations) that suggest that the F, LB and B water molecules of fluorine correspond to the LB and F water molecules of the sulfonate groups.

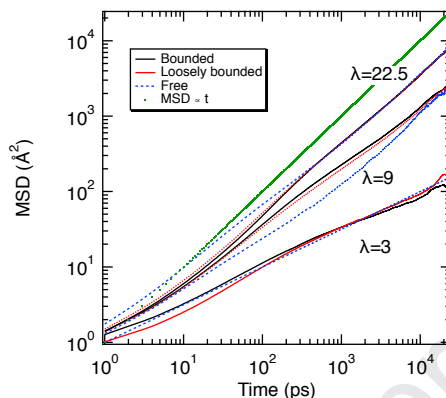


Figure 10: Mean square displacements of water in the LB region around F1 for $\lambda = 3, 9$ and 22.5.

This result, based on the position of the first minimum in RDF, suggests
 340 that the water molecules could form a conducting corridor between sulfonate
 groups as illustrated in Figure 11a. By calculating the number of hydrogen
 bonds [48] of water molecules we highlighted that the water corridor was based
 on the hydrogen bonds between molecules forming a single file structure (see
 Figure 11b). Interestingly, this structure is well known to involve a sub-diffusive
 345 dynamics [54] that is in line with the value of $\langle \alpha \rangle$ reported in Figure 4. This
 water corridor structure could be then a conducting pathway of protons in the
 Grotthuss mechanism [13]. Although classical molecular dynamics simulations
 can not take into account the transfer of the proton between water molecules,
 the existence of water corridors between sulfonate groups where each molecule
 350 are connected by the hydrogen bonds will be in favor of the transport of proton
 based on the Grotthuss process.

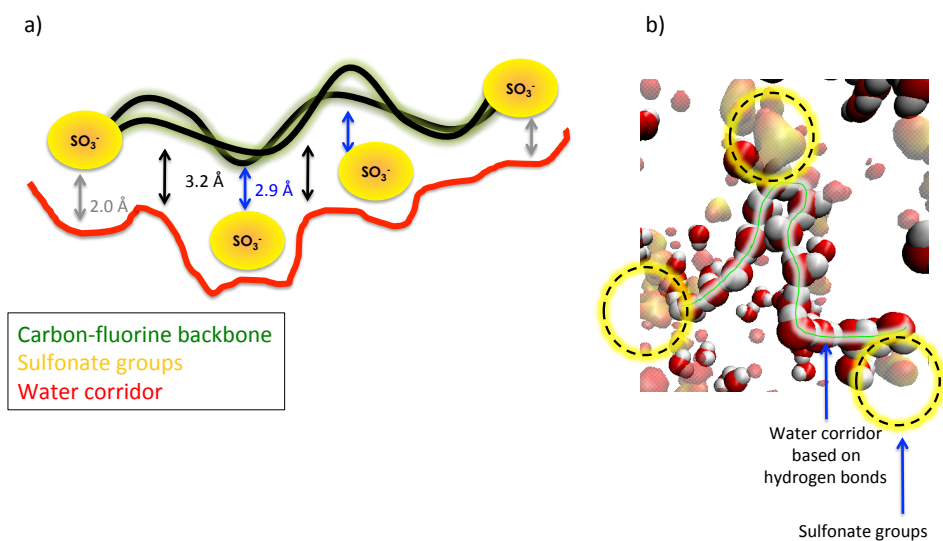


Figure 11: a) Illustration of a water corridor along the carbon-fluorine backbone b) Snapshot at $\lambda = 7$ highlighting a water corridor based on the hydrogen bonds between water molecules. Water molecules have been chosen by considering the hydrogen bonds number using the geometric criteria defined by Luzar and Chandler [48].

Indeed, as shown in Figure 12a the MSD of hydronium ions are around 40-60 \AA^2 whatever λ i.e. a displacement of 6-7 \AA involving a very strong anchoring around the sulfonate groups. At $\lambda=7$, the MSD is 18.1 \AA^2 while at $\lambda=11.5$ the MSD is 13.5 \AA^2 . The difference between both water uptake corresponds to 4.6 \AA^2 i.e. a displacement of 2 \AA along the sulfonate groups. As this difference is not really significant the MSD of ions can be then considered as similar between $\lambda=7$ and $\lambda=11.5$. Figure 12b which report the RDF between hydronium ions and the sulfonate groups evidences several peaks between 2 and 6 \AA that corroborate a strong interactions. This result is in good agreement with the work of Devanathan and coworkers that has shown a very slow proton dynamics in line with the neutron scattering experiments. [29]

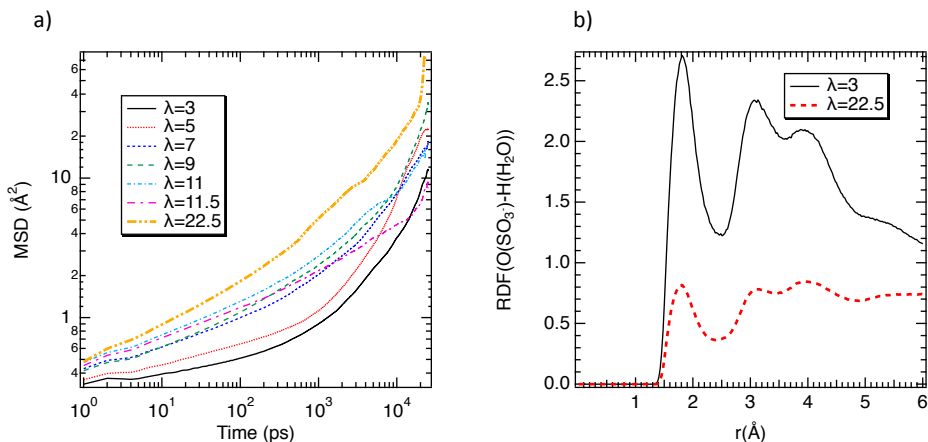


Figure 12: a) Mean square displacements of H_3O^+ for all uptakes. b) RDF between oxygen atoms of SO_3^- and hydrogen atoms of H_3O^+ .

The hydronium ions located beyond 6 Å in the second solvation shell of the sulfonate groups correspond probably to the ions located in the first shell of another SO_3^- . That is probably the result of the polymer folding. To ensure this hypothesis we report in Figure 13a the RDF between SO_3^- . Figure 13a shows three maxima at 2.5 Å, 3.0 Å and 5.0 Å that highlight an interdigitated structure. At high water uptake, Figure 13b exhibits the self-organizing of Nafion leading to a phase separation between water and the Nafion membrane. Interestingly, as shown in Figure 13b the hydrophilic sulfonate groups point toward the water phase involving a Nafion surface decorated with SO_3^- interacting with water molecules. Therefore a water/Nafion interface is evidenced with a surface rich in SO_3^- such that water molecules could jump between sulfonate sites. This result is in good agreement with the work of G.A. Voth and C.K. Knox because their very large scale simulations revealed fast intercluster 'bridge' formation and network percolation. Sulfonate groups were found inside these bridges and played a significant role in percolation. Sulfonates also strongly aggregated around and inside clusters leading to the hydrophilic-hydrophobic interface. [55]

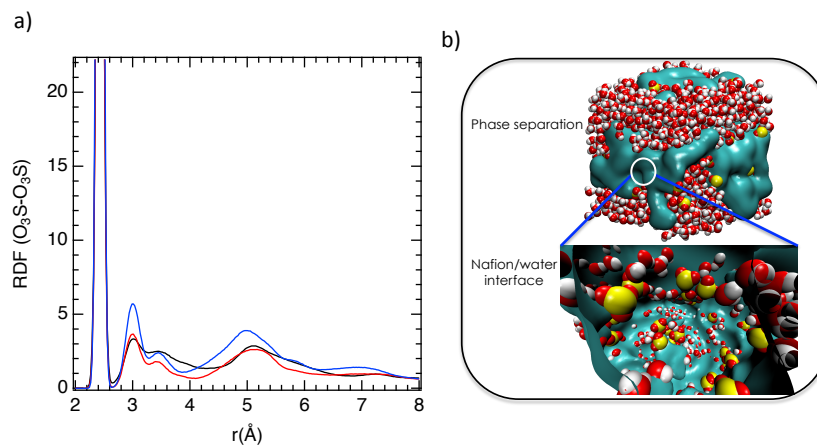


Figure 13: a) RDF between sulfur atoms of SO_3^- at $\lambda=3$. b) Snapshots of hydrated Nafion and the water/Nafion interface at $\lambda = 22.5$.

380 Translational jumps were quantified by calculating the self Van Hove function ($G_S(r,t)$) where r and t are the position and time of guest molecules, respectively, $G_S(r,t) = \frac{1}{N} \left\langle \sum_{j=1}^N \delta[\mathbf{r} - (\mathbf{r}_j(t) - \mathbf{r}(0))] \right\rangle$, $r_j(t)$ is the atomic position of j particle at time t and N is the number of water molecules. Figure 14a reports $G_S(r,t)$ from 1 ps to 300 ps for a loading of $\lambda = 3$. The maximum of the
 385 Van Hove function remains at a distance lower than 1.5 \AA from 2 ps to 300 ps, which indicates small displacements and an absence of translational jumps in this time interval. However a growth of a shoulder is observed around 3.0 \AA that could correspond to the translational jumps between two sulfonate groups side by side.

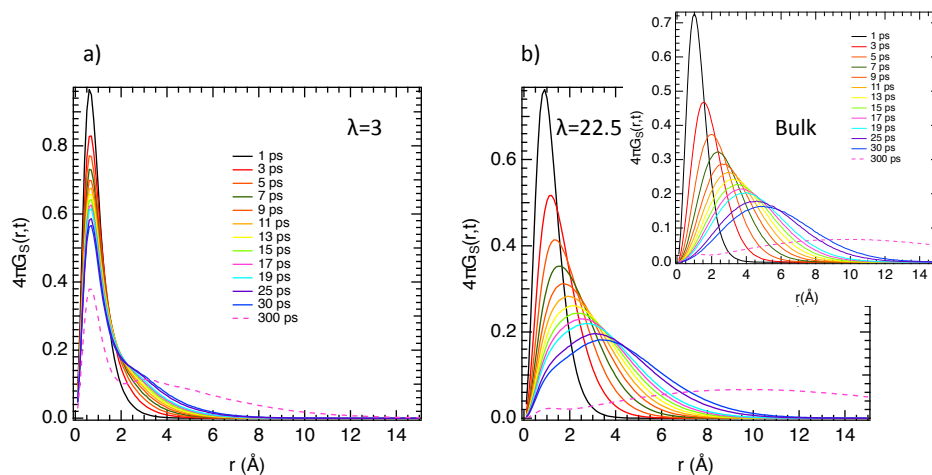


Figure 14: Self van Hove functions as a function of the covered distance for some times for $\lambda=3$ a), 22.5 b) and for the water liquid bulk phase (inset of part b))

390 4. Concluding Remarks

In this work diffusion of confined water into the Nafion membrane was examined.

A sub-diffusive regime was evidenced whatever the hydration level. This regime was important at low λ whereas it was found decreasing in favor of
 395 diffusive one as the hydration rate increased. A first progressive sub-diffusive to diffusive transition as a function of the water uptake was then evidenced. Structural analyses were then carried out to structurally understand this dynamical behavior. Cluster size study evidenced a polydisperse distribution of nanometric cluster inducing nanophases. The first hypothesis was that, at low
 400 water uptake, water molecules are trapped in these nanophases and locally diffusive in small size environment inducing a sub-diffusive regime. The percolation occurs from higher hydration level such that all nanophases begin to be gathered in a single interconnected structure. The progressive transition between sub-diffusive to diffusive regime as the increase in water uptake was then, in
 405 first time, ascribed to the structural process. More especially, we have shown

that the water molecules firstly adsorb on the sulfonate groups and grow to form these water nanophases around sulfonate groups. As well water uptake increases nanophases grows to finally interconnect and to form a single structure.

However, we have shown the presence of a permanent single interconnected
410 water structure involving that the sub-diffusive to diffusive transition could be only the result of the time assembly of nanophases. Analyses of the local dynamics helped us to examine the molecular origin of this dynamics. From radial distribution calculations three regions around the sulfonate groups and the fluorine atoms were drawn, corresponding to the bounded (B) water, to
415 the loosely bounded (LB) water and the free (F) water. Local dynamics have been then investigated by considering this spatial division. In the (B) boundary of sulfonate groups the sub-diffusive regime was always evidenced and was found progressively timely replaced by the diffusive one in all range of λ . The so-observed sub-diffusive regime was ascribed to the water molecules strongly
420 anchored close to the sulfonate groups and locally diffusive in small size environment. We evidenced that the sub-diffusive to diffusive time transition was the result of the elapsed time before to escape of the attractive interactions of the sulfonate groups. This phenomenon was found independent of the water uptake. The diffusive regime is then possible from the diffusive corridors induced
425 by the percolation of the hydrogen bonding network. On the contrary in both LB and F regions the dynamics were found rather diffusive. As a function of the location of the water molecules their dynamics can be then drastically different involving dynamical heterogeneities.

The progressive transition between sub-diffusive to diffusive regime as a function
430 of the hydration level is ascribed to the respective weight of diffusive dynamics that increase with respect to the sub-diffusive regime due to the increase in diffusive pathways at the expense of the localized dynamics.

Close to the fluorocarbon backbones the dynamics of water molecules is also sub-diffusive but faster than the water located close to the sulfonate groups.
435 Additionally, time dynamical transition was not observed. Dynamics of water molecules depends on their positions into the membranes (anchored in the

vicinity of sulfonate groups, close to the fluorocarbon backbones and in the nanometric channels) involving then dynamical heterogeneities. Furthermore we highlighted the existence of water corridors based on the hydrogen bonds
440 between molecules leading to the formation of a single file structure which is in line with a sub-diffusive dynamics. This water corridor structure is assimilated to a possible conducting pathway of protons in a Grotthuss mechanism. From a Nafion standpoint, we exhibited an interdigitated structure where the sulfonate groups are interleaved. At high water uptake, we exhibit the self-organizing
445 of Nafion leading to a phase separation between water and the Nafion membrane. Eventually, a specific interfacial organization was evidenced involving a hydrophilic surface and then a water/Nafion interface. At the microscopic scale, water molecules diffuse then close to the sulfonate groups (hydrophilic regions) and slide along the Fluorine groups (hydrophobic zones). However, at
450 the macroscopic scale (corresponding to the NMR results) the water transport corresponds to the diffusion through the hydrophilic nanophases defined by the sulfonate groups leading to a hydrophilic surface. That suggests then an interfacial transport at the membrane/fluid interface.

Competing Financial Interest

455 The authors declare no competing financial interests.

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465 **References**

- [1] I. EG&G Technical Services, Fuel Cell Handbook, U.S. Department of Energy - Office of Fossil Energy - National Energy Technology Laboratory, P.O Box 880 - Morgantown, West Virginia 26507-0880, 2004.
- [2] W. Goddard–III, B. Merinov, A. V. duin, T. Jacob, M. Blanco, V. Molinero, S. Jang, Y. Jang, Multi-paradigm multi-scale simulations for fuel cell catalysts and membranes, *Mol. Sim.* 32 (2006) 251.
- [3] S. Cui, J. Liu, M.-E. Selvan, D. Keffer, B. Edwards, W. Steele, A molecular dynamics study of a nafion polyelectrolyte membrane and the aqueous phase structure for proton transport, *J. Phys. Chem. B* 111 (2007) 2208.
- 475 [4] P. Laflamme, A. Beaudoin, T. Chapaton, C. Spino, A. Soldera, Simulated infrared spectra of triflic acid during proton dissociation, *J. Comp. Chem.* 33 (12) (2012) 1190.
- [5] P. Laflamme, A. Beaudoin, T. Chapaton, T. Spino, A. Soldera, Molecular modeling assisted design of new monomers utilized in fuel cell proton exchange membranes, *J. Mem. Sci.* 401 (2012) 56.
- 480 [6] N. Metatla, S. Palato, A. Soldera, Change in morphology of fuel cell membranes under shearing, *Soft Matter* 9 (2013) 11093.
- [7] H. Mendil–Jakani, S. Pouget, G. Gebel, P. Pintauro, Insight into the multiscale structure of prestretched recast nafion membranes: Focus on the crystallinity features, *Polymer* 63 (2015) 99.
- 485 [8] A. Fleury, A. Godey, F. Laflamme, A. Ghoufi, A. Soldera, Is fine-grained simulation able to propose new polyelectrolyte membranes?, *Fuell Cells* 16 (2016) 675.
- [9] Q. Berrod, S. Hanot, A. Guillermo, S. Mossa, S. Lyonard, Water sub-diffusion in membranes for fuel cells, *Scientific Reports* 7 (2017) 8326.
- 490

- [10] N. Martinez, A. Morin, Q. Berrod, B. Frick, J. Olivier, L. Porcar, G. Gebel, S. Lyonnard, Multiscale water dynamics in a fuel cell by operando quasi elastic neutron scattering, *J. Phys. Chem. C* 122 (2018) 1103.
- [11] M. Tripathy, P. S. Kumar, A. Deshpande, Molecular structuring and percolation transition in hydrated sulfonated poly(ether ether ketone) membranes, *J. Phys. Chem. B* 121 (2017) 4873.
- [12] A. Vishnyakov, R. Mao, M.-T. Lee, A. Neimark, Coarse-grained model of nanoscale segregation, water diffusion, and proton transport in nafion membranes, *J. Chem. Phys* 148 (2018) 024108.
- [13] C. de Grotthuss, Sur la décomposition de l'eau et des corps qu'elle tient en dissolution à l'aide de l'électricité galvanique, *Ann. Chim.* 58 (1806) 54.
- [14] T. Peckham, S. Holderoft, Structure-morphology-property relationships of non-perfluorinated proton-conducting membranes, *Adv. Mater.* 22 (2010) 4667.
- [15] Z. Lu, G. Polzos, D. Macdonald, E. Manias, State of water in perfluorosulfonic ionomer (nafion 117) proton exchange membranes, *J. Elec. Soc.* 155 (2008) B163.
- [16] F. Xu, S. Leclerc, D. Canet, Nmr relaxometry study of the interaction of water with a nafion membrane under acid, sodium, and potassium forms. evidence of two types of bound water, *J. Phys. Chem. B* 117 (2013) 6534.
- [17] T. Shimoaka, C. Wakai, T. Sakabe, S. Yamazaki, T. Hasegawa, Hydration structure of strongly bound water on the sulfonic acid group in a nafion membrane studied by infrared spectroscopy and quantum chemical calculation, *Phys. Chem. Chem. Phys.* 17 (2015) 8843.
- [18] T. Zawodzinski, M. Neeman, L. Sillerud, S. Gottesfeld, Determination of water diffusion coefficients in perfluorosulfonate ionomeric membranes, *J. Phys. Chem.* 95 (1991) 6040.

- [19] S. Motupally, A. Becker, J. Weidner, Diffusion of water in nafion 115 membranes, *J. Elec. Soc.* 147 (2000) 3171.
- 520 [20] S. Ochi, O. Kamishima, J. Mizusaki, J. Kawamura, Investigation of proton diffusion in nafion 117 membrane by electrical conductivity and nmr, *Solid State Ionics* 180 (2009) 580.
- [21] D. Hallinan, M. D. Angelis, M. Baschetti, G. Sarti, Y. A. Alabd, Non-fickian diffusion of water in nafion, *Macromolecules* 43 (2010) 4667.
- 525 [22] Q. Zhao, P. Majsztrik, J. Benziger, Diffusion and interfacial transport of water in nafion, *J. Phys. Chem. B* 115 (2010) 2717.
- [23] Q. Duan, H. Wang, J. Benziger, Transport of liquid water through nafion membranes, *J. Mem. Sci.* 88 (2012) 392.
- [24] O. Sel, L. T. T. Kim, C. Debiemme-Chouvy, C. Gabrielli, C. Laberty-Robert, H. Perrot, Determination of the diffusion coefficient of protons in nafion thin films by ac-electrogravimetry, *Langmuir* 29 (2013) 13655.
- 530 [25] J. Song, O. Han, S. Han, Nanometer-scale water and proton diffusion heterogeneities across water channels in polymer electrolyte membranes, *Angew. Chem. Int. Ed.* 54 (2015) 3615.
- 535 [26] Z.-Z. Li, L. Chen, W.-Q. Tao, Molecular dynamics simulation of water permeation through the nafion membrane, *Numerical heat transfer, Part a* 70 (2016) 1232.
- [27] C. Casieri, A. Monaco, F. D. Luca, Evidence of temperature–induced sub-diffusion of water on the micrometer scale in a nafion membrane, *Macromolecules* 43 (2010) 638.
- 540 [28] T. Mabuchi, T. Tokumasu, Relationship between proton transport and morphology of perfluorosulfonic acid membranes: A reactive molecular dynamics approach, *J. Phys. Chem. B* 122 (2018) 5922.

- [29] R. Devanathan, A. Venkatnathan, M. Dupuis, Atomistic simulation of
545 nafion membrane: I effect of hydration on membrane nanostructure, *J. Phys. Chem. B* 111 (28) (2007) 8069.
- [30] D. Theodorou, U. Suter, Detailed molecular structure of a vinyl polymer
glass, *Macromolecules* 18 (1985) 1467.
- [31] D. Hofmann, L. Fritz, J. Ulbrich, C. Schepers, M. Bohning, Detailed-
550 atomistic molecular modeling of small molecule diffusion and solution processes in polymeric membrane materials, *Macromol. Theory Simul.* 9 (2000) 293.
- [32] A. Ghoufi, G. Maurin, Hybrid monte carlo simulations combined with a
phase mixture model to predict the structural transitions of a porous metal-
555 organic framework material upon adsorption of guest molecules, *J. Phys. Chem. C* 114 (2010) 6496.
- [33] S. L. Mayo, B. D. Olafson, W. Goddard—III, Dreiding: a generic force field
for molecular simulations, *J. Phys. Chem.* 94 (1990) 8897.
- [34] S. Jang, V. Molinero, T. Cagm, W. Goddard—III, Nanophase-segregation
560 and transport in nafion 117 from molecular dynamics simulations: Effect of monomeric sequence, *J. Phys. Chem. B* 108 (2004) 3149.
- [35] M. Baaden, M. Burgard, G. Wipff, Tbp at the water-oil interface: The
effect of tbp concentration and water acidity investigated by molecular
dynamics simulations, *J. Phys. Chem. B* 105 (2001) 11131.
- 565 [36] Y. Wu, H. Tepper, G. Voth, Flexible simple point-charge water model with improved liquid-state properties, *J. Chem. Phys* 124 (2006) 024503.
- [37] J. Abascal, C. Vega, A general purpose model for the condensed phases of
water: Tip4p/2005, *J. Chem. Phys.* 123 (2005) 23505.
- [38] P. Ewald, Die berechnung optischer und elektrostatischer gitterpotentiale,
570 *Ann. Phys.* 64 (1921) 253.

- [39] T. Darden, L. Perera, L. Pedersen, New tricks for modelers from the crystallography toolkit: the particle mesh ewald algorithm and its use in nucleic acid simulations, *Structure* 7 (1999) R55.
- [40] I. Todorov, W. Smith, K. Trachenko, M. Dove, Dlpoly3: new dimensions
575 in molecular dynamis simulations via massive parallelism, *J. Mater. Chem.* 16 (2006) 1911.
- [41] W. Hoover, Canonical dynamics: Equilibrium phase-space distributions, *Phys. Rev. A* 31 (1985) 1695.
- [42] A. Ghoufi, P. Malfreyt, Local description of surface tension through ther-
580 modynamic and mechanical definitions, *Mol. Sim.* 39 (2012) 603.
- [43] F. Biscay, A. Ghoufi, P. Malfreyt, Surface tension of water-alcohol mixtures from monte carlo simulations, *J. Chem. Phys* 134 (2011) 044709.
- [44] J.-P. Bouchaud, A. Georges, Anomalous diffusion in disordered media: Sta-
585 tistical mechanisms, models and physical applications, *Physics Reports* 195 (4) (1990) 127.
- [45] E. Flenner, J. Das, M. Rheninstadter, I. Kosztin, Subdiffusion and lateral diffusion coefficient of lipid atoms and molecules in phospholipid bilayers., *Phys. Rev. E* 79 (2009) 011907.
- [46] J. Perrin, S. Lyonnard, F. Volino, Quasielastic neutron scattering study
590 of water dynamics in hydrated nafion membranes, *J. Phys. Chem. C* 111 (2007) 3393.
- [47] S. Stoddard, Identifying clusters in computer experiments on systems of particles, *J. Comp. Physics* 27 (2) (1978) 291.
- [48] A. Luzar, D. Chandler, Effect of environnement on hydrogen bond dynam-
595 ics in liquid water, *Phys. Rev. letters* 76 (1996) 928.
- [49] W. Hsu, T. D. Gierke, Ion transport and clustering in nafion* prefluorinated membranes, *J. Mem. Sci.* 13 (3) (1983) 307.

- [50] S. Dixit, J. Crain, W. poon, J. Finey, A. Soper, Molecular segregation observed in a concentrated alcohol-water solution, *Nature* 416 (2002) 829.
- 600 [51] X. Wu, X. Wang, G. He, J. Benziger, Differences in water sorption and proton conductivity between nafion and speek, *Pol. Phy. part B* 49 (2011) 1437.
- [52] A. Bizzarri, S. Cannistraro, Molecular dynamics of water at the protein-solvent interface, *J. Phys. Chem. B* 106 (2002) 6617.
- 605 [53] F. Goujon, P. Malfreyt, J. Simon, A. Boutin, B. Rousseau, A. Fuchs, Monte carlo versus molecular dynamics simulations in heterogenous systems: An application to the n-pentane liquid-vapor interface, *J. Chem. Phys.* 121 (2004) 12559.
- [54] H. Jobic, Observation of single-file diffusion in a mof, *Phys. Chem. Chem. Phys* 18 (2016) 17190.
- 610 [55] C. Knox, G. Voth, Probing selected morphological models of hydrated nafion using large-scale molecular dynamics simulations, *J. Phys. Chem. B* 114 (9) (2010) 3205.

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A.G, B.G., F.G. and A.F. performed all Molecular Dynamics simulations. All authors were involved in the interpretation of water transport. The manuscript was written by A.G and A.S.

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Declaration of interests

x The authors declare that they have no known competing financial interests or personal relationships that could have appeared to influence the work reported in this paper.

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