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3 **Improvement of water resistance of hemp woody substrates through deposition of**
4 **functionalised silica hydrophobic coating, whilst retaining excellent moisture**
5 **buffering properties**
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13 *Pretot⁵, Atif Hussain¹, Martin Ansell¹, Michael Lawrence¹*
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40 **Corresponding author; e-mail: y.jiang@bath.ac.uk/yunhongjiang@yahoo.com*
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44 **Abstract**
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46 This paper reports on the development of a novel treatment for hemp shiv that
47 improves resistance to liquid water and protects hemp shiv from bio-degradation
48 without impacting the natural ability of the shiv to buffer moisture vapour. The
49 hydrophobic surface was produced by depositing silica nanoparticles, which were
50 prepared by the functionalisation of silica nanoparticles with hexamethyldisilazane
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3 (HMDS). The surface chemical composition was determined showing replacement of
4 surface silanols on the silica nanoparticles with trimethyl groups. The specific
5 surface area of silica nanoparticles decreased after the trimethylsilyl treatment. The
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60 surface area of silica nanoparticles decreased after the trimethylsilyl treatment. The

surface modifications reduced the level of hysteresis between absorption and desorption isotherms and also the total amount of moisture absorbed by the silica nanoparticles. The surface of the hemp shiv was initially hydrophilic, but became hydrophobic once the material was treated, demonstrating contact angle with water of 120°. The results show the coating layer of functionalised silica on the hemp shiv reduced water absorption from 400% (untreated shiv) to 250%. However, the moisture buffer value results showed that the coating films do not limit the access of moisture to adjacent pores in the hemp shiv and the functionalised silica coating layer retains the moisture buffering ability of hemp shiv.

33 **Keywords**

34 Hydrophobic coating; Nanostructure; Moisture adsorption; Hexamethyldisilazane;
35 Renewable resources; Moisture buffer value
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Introduction

The construction industry, in common with industry in general, now has a much greater awareness of environmental and sustainability issues, and the drive from consumers to reduce the reliance on fossil based materials. Considerable attention has been focused on the utilisation of bio-based materials for construction to offer a number of benefits including thermal¹, hygroscopic² and acoustic³ properties, whilst simultaneously delivering renewable and sustainable solutions to reduce detrimental environmental effects. Apart from timber, other bio-based materials have shown promise as construction materials, including hemp shiv, formed from the core of the hemp plant (*Cannabis Sativa L.*), since it has the ability to naturally regulate the indoor environment through a reduction in humidity variations as the material attracts, holds and releases moisture in response to changes in ambient conditions. This ability is due to the highly porous nature of hemp shiv and its hydrophilic character since it contains up to 44% cellulose^{2, 4-6}. These characteristics mean that hemp shiv is greatly influenced by the environment and is susceptible to decay if exposed to unsuitable environmental conditions. Therefore, it is a crucial step to modify the surface of hemp shiv in order to avoiding the problem of decay and to increase durability without detrimentally affecting the desirable moisture buffering characteristics.

One technique is to coat or chemically modify the hydrophilic surface to render it hydrophobic⁷⁻¹¹. Approaches have been reported in the literature where nanotechnology methods have been employed to prepare multifunctional materials

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3 which are then deposited as coatings on plant materials¹²⁻¹⁴. This usually involves a
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5 sol-gel process to create inorganic or inorganic-organic hybrid coatings which can
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7 render a (hydrophilic) cellulose based material hydrophobic. Water sorption sites are
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9 reduced by blocking the hydroxyl groups of the cell wall through steric or chemical
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11 modification¹⁵⁻²⁰. The generic chemical reactions underpinning sol-gel chemistry are
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13 well known. In these examples the assumption is that the precursor is an alkoxide,
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15 such as a silane. During hydrolysis and condensation reactions, siloxane networks
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17 are formed by precursors such as tetraethoxy silane (TEOS). These reactions provide
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19 the structural backbone of sol-gel systems. The structural evolution of sol-gel is quite
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21 complex. The different chemical paths the materials follow are influenced by acidic
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23 or basic catalysis, which has been well described by Bergna²¹. Significant effort has
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25 been devoted to preparing organically modified silicate materials with improved
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27 hydrophobic properties to modify the hydrophilic character of silica based materials
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29 and coatings. Surface modification of silica nanoparticles to achieve hydrophobic
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31 behaviour has been widely reported. A common route is to graft trimethyl groups to
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33 the surface via silylation reactions that replace the hydrophilic surface silanols²²⁻²³.
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35 Applications of sol-gel based materials have been reported to improve water,
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37 chemical, fire and weather resistance to wood, and also confer antibacterial
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39 properties²⁴⁻²⁷. Liu et al. reported a superhydrophobic coating on the wood surface
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41 via a solution-immersion process¹⁵. A sol-gel method using methyltrimethoxysilane
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43 and hexadecyltrimethoxysilane was used to add a hydrophobic polysiloxane coating
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45 to wood reported by Tshabalala et al²⁸. Tang et al. assembled silica and silver
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47 nanoparticles on the surface of wool fabric to produce super-hydrophobic property
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3 and antibacterial activity²⁹. The application of sol-gel based hydrophobic coatings to
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5 bio-based materials (wood, cotton, etc.) has been discussed in the literature, but there
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7 is a general lack of information on the effect on the moisture buffering capabilities of
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9 these materials. However, the application of hydrophobic Carnauba wax particles as
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11 a non-continuous coating was reported to maintain the moisture buffering efficiency
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13 of wood by Lozhechnikova et al³⁰⁻³¹.
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19 The aim of this work was to develop a modification technique that increases the
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21 resistance of hemp shiv to liquid water while improving the chemical compatibility
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23 of the hemp shiv with the binder without reducing the natural ability of hemp shiv
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25 to buffer moisture vapour. This is because the presence of natural sugars on the
26
27 hemp shiv surface inhibit the setting of hydraulic binders. As a first step SiO₂
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29 nanoparticles were prepared using the Stöber method. These colloidal dispersions of
30
31 silica in alcohol were then functionalised to give hydrophobic behaviour. The hemp
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33 shiv were then immersed in the suspension of functionalised silica particles,
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35 withdrawn and then dried to leave a thin film of the particles on the shiv. The
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37 hydrophobicity and moisture buffer properties of the treated hemp shiv were
38
39 studied by sessile drop analysis and the Nordtest test³².
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47 **Experimental**

48 **Materials**

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50 The synthesis of silica nanoparticles was carried out by using tetraethoxysilane
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52 (TEOS) (Silanes & Silicones, Stockport, UK) as precursor, ammonia, industrial
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3 methylated spirit (IMS, 99% ethanol, 1% methanol) (Sigma-Aldrich, Gillingham, UK),
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5 and water which was deionised in-house. The functionalising agent,
6
7 hexamethyldisilazane, was also purchased from Silanes & Silicones. The hemp shiv
8
9 were provided by Cavac Biomateriaux (Cavac, Sainte-Gemme-La-Plaine, France).
10
11 The chemical and physical characteristics of the hemp shiv materials used are as
12
13 reported by Jiang et al³³.
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19 Synthesis and modification of SiO₂ NPs

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21 The Stöber process was followed in order to synthesise silica nanoparticles, at room
22
23 temperature³⁴. First, a solution containing appropriate quantities of IMS, ammonia
24
25 and deionised water was stirred for 5 minutes to ensure complete mixing. Separately
26
27 the TEOS was mixed with IMS and the mixture was then added to the first solution.
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29 The reaction proceeded at ambient temperature for 48 h according to concentrations
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31 of reactants, similarly to the protocol followed by Bourebrab et al.¹⁹. The formulation
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33 reached 4.3wt% silica on completion of the reaction. The silica nanoparticles are
34
35 referred to as TSS4.
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42 To functionalise the silica a reaction with hexamethyldisilazane was undertaken in
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44 the parent suspension by heating at 65 °C for 18 h so that the silanols present on the
45
46 surface of the silica nanoparticles would react with the silazane by condensation. The
47
48 functionalised silica is referred to as TSS4T1.
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3 To prepare for the hydrophobic treatment of the hemp shiv, the hemp shiv was oven
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5 dried at 90°C for one hour prior to coating. The shiv was then immersed in the
6
7 TSS4T1 aqueous suspension for 1 min with constant mixing. The coated hemp shiv
8
9 was then shaken to remove any visible liquid which would have remained on the
10
11 surface of the shiv, and then the bio-materials were oven dried at 90°C for one hour.
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14 15 16 17 Sessile drop assessment

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19 The static sessile drop method was used to determine wetting behaviour of the
20
21 samples. The materials and methods used are as reported by Bourebrab et al¹⁹. Three
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23 to ten droplets of deionised water of approximately 2 µL were placed onto the
24
25 samples. An image of the droplets was captured after waiting one or two seconds for
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27 the drop to reach equilibrium before taking a measurement, thus allowing key
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29 dimensions to be measured.
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35 36 Materials Characterisation

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38 The morphology of the surface coating was characterised using a scanning electron
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40 microscope (SEM) (JEOL SEM-6480LV, Tokyo, Japan) and a field emission scanning
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42 electron microscope (FE-SEM) (JEOL FESEM6301F). The materials and methods used
43
44 are as reported by Jiang et al³³. TSS4T1 were also deposited on copper grids for
45
46 transmission electron microscopy (TEM) (JEOL-2100 Plus) observation, and the
47
48 accelerating voltage was 200 kV. The thickness of coating layers and the internal
49
50 surface of hemp shiv were characterised by a dual beam FEI Quanta 3D scanning
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52 electron microscope (FIB-SEM) with a platinum deposition. The surface areas and
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3 the pore size distributions of the samples were calculated from nitrogen adsorption-
4 desorption isotherms at 77 K using the BET method (3Flex, UK). Size measurement
5 of the silica particles was performed using dynamic light scattering (DLS) on a
6 Malvern Zetasizer Nano-S (Malvern Instruments, Malvern, UK), at room
7 temperature, via a back-scattering method. FTIR analysis of treated and untreated
8 hemp shiv was carried out by using a PerkinElmer FTIR spectrometer model
9 Frontier. Transmittance spectra were recorded in an Attenuated Total Reflection
10 (ATR) mode and collected with 2 cm^{-1} resolution, and 10 scans were accumulated for
11 each spectrum in the range $4000\text{-}600\text{ cm}^{-1}$. For this analysis the silica suspension was
12 dried, and crushed to break down the aggregates. The same amount of sample was
13 ground with a 1:300 dilution of KBr to produce pellets. Transmittance spectra were
14 recorded in transmission and collected with 2 cm^{-1} resolution, and 10 scans were
15 accumulated for each spectrum in the range $4000\text{-}600\text{ cm}^{-1}$.

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36 Study of adsorption and desorption behaviour of water vapour on TSS4 and TSS4T1

37 The adsorption and desorption of water vapour were investigated on TSS4 and
38 TSS4T1 using a Surface Measurement Systems dynamic vapour sorption apparatus
39 (DVS, Surface Measurement Systems, UK). The materials and methods used are as
40 reported by Hill et al³⁵.

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49 Water absorption test

50 The hemp shiv samples were placed in water at room temperature without applying
51 any additional force to immerse them. The hemp shiv pieces floated to the surface as
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3 their density was lower than water. Therefore, the majority of the water absorption
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5 took place through the capillary behaviour in the hemp shiv samples. The samples
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7 were taken out of water at specified time intervals, and the mass readings were
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9 taken within 30 seconds of removal using a balance with an accuracy of ± 0.1 mg.
10
11 Any water visible on the surface of hemp shiv samples was shaken off before the
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13 measurement. The water absorption was calculated by the mass change (%) between
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15 the initial mass and the mass after 24 hours of water absorption. The measurements
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17 were performed in triplicate and the average reading was reported³⁶.
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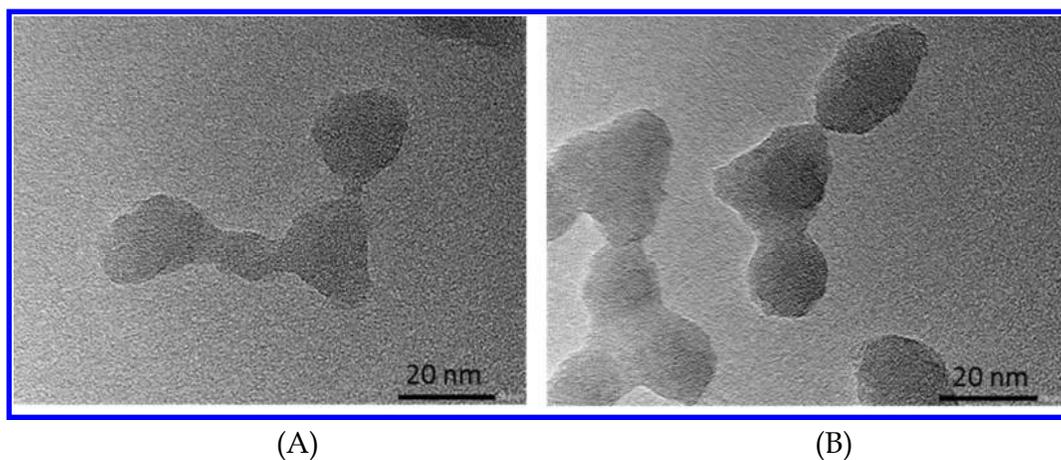
24 Moisture buffering analysis

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26 In order to evaluate the effectiveness of moisture buffering, the Practical Moisture
27
28 Buffer Value (MBV) in $\text{g/m}^2.\%RH$ was calculated. The moisture buffering
29
30 experiments were performed in accordance with the Nordtest method³². MBV
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32 defines the amount of moisture transported into or out of a material per unit of open
33
34 surface area, during a specified period of time, when the material is exposed to cyclic
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36 variations in relative humidity (8 hours at 75% RH, followed by 16 hours at 33 %RH).
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38 Samples were put in watertight cylindrical containers, 10 cm in diameter, open at the
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40 top. Containers were chosen to be higher than the penetration depth. Full details of
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42 the materials and methods used are as reported by Collet et al³⁷.
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49 **Results and Discussion**

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51 Characterisation of the silica NPs and the functionalised silica NPs
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3 TEM images of the Stöber silica before and after functionalisation are shown in Fig 1.
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5 Fig. 1A shows the TSS4 nanoparticles had formed irregular agglomerations.
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7 Typically, the size of the TSS4 particle is between 15-30 nm. Fig. 1B shows that the
8
9 TSS4T1 nanoparticles have a similar particle size to the untreated primary silica
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11 nanoparticles (TSS4). No distinct coating layers from the trimethylsilylation were
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13 observed on the outside of the TSS4T1 nanoparticles because the silylation will
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15 generate a surface assembled monolayer, which is likely to be less than 1 nm. Overall,
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17 silica nanoparticles did not exhibit notable changes in appearance after modification.
18
19 The particles are spheroidal with both major and minor axis diameters in the range
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21 of 15-30 nm.
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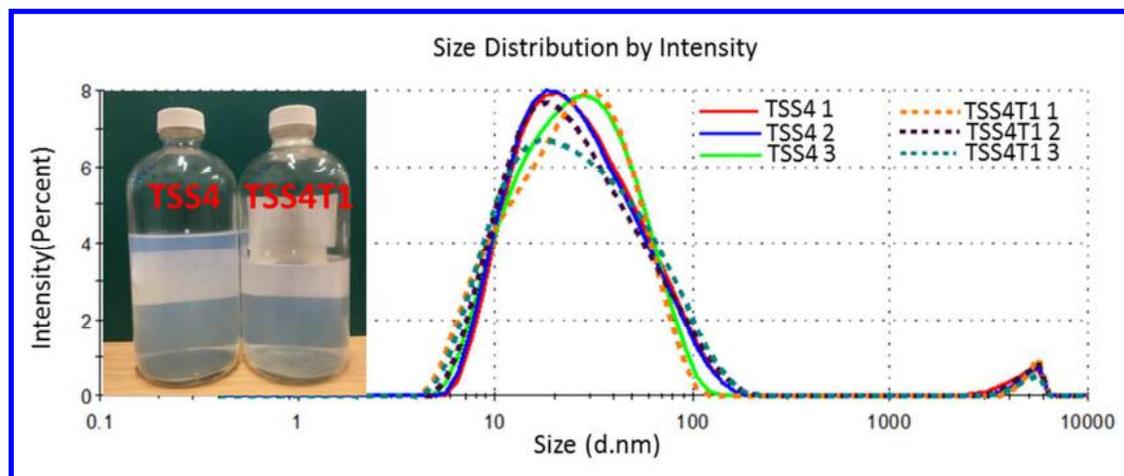
Figure 1: TEM images of the TSS4 (A) and the TSS4T1 (B)

Suspensions of both TSS4 and TSS4T1 nanoparticles were analysed by Dynamic Light Scattering (DLS), in order to provide estimates of their hydrodynamic sizes. The average measured diameters for the TSS4 and TSS4T1 nanoparticles were 22.6 and 20.8 nm, respectively, in Table 1. Fig. 2 also shows that the two suspensions had

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3 similar size distributions. The size distribution of silica nanoparticles shifted to
4 slightly smaller values after surface modification. The corresponding polydispersity
5 indexes (PDI) for DLS measurement are also reported in Table 1. Stetefeld et al
6 reported that a PDI of ~ 0.1 indicates a monodisperse system³⁸. This indicates that
7 both suspensions are not monodispersed and there are some aggregates around a
8 few micrometres present in both suspensions of TSS4 and TSS4T1, as shown in Fig. 2.
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Table 1. The z-average diameters and polydispersity index (PDI) obtained by Zetasizer Nano

Samples	T (°C)	Z-Ave (nm)	Std Dev	PDI	Std Dev
TSS4	25	22.6	± 0.17	0.3	± 0.002
TSS4T1	25	20.8	± 0.39	0.4	± 0.013



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3 Figure 2: *The intensity size distribution of the silica NPs and the functionalised silica NPs*
4 *versus particle diameter in nanometres. The different colours represent repeat measurements*
5 *of the same sample. The inserts are the photo images of the sample suspension.*
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12 The specific surface areas and the parameter C values of the BET equation for TSS4
13 and TSS4T1 are shown in Fig. 3 and Table 2. They both show typical type IV
14 isotherms according to the IUPAC classification³⁹⁻⁴⁰, which is characteristic of
15 mesoporous materials. Both the values decreased after the surface modification.
16 Since the parameter C is related to the affinity of the solid with the adsorbate (the N₂
17 molecules) and so to the heat of adsorption. The higher the value of C, the higher the
18 interaction between adsorbate and surface of sample⁴¹. TSS4T1 has a lower value for
19 C (38) in the BET equation and TSS4 has a higher value for C (143), indicating a
20 reduction in the strength of the interaction between the nitrogen molecules and the
21 molecular surface. This can be explained by the ability of the TMS groups to change
22 the silica surface to a nonpolar surface from a polar surface⁴². The pore size
23 distributions of the samples before and after the modification are shown in Fig.3 and
24 Table 2. The increase in pore diameter as a result of the surface modification is 0.74
25 nm.
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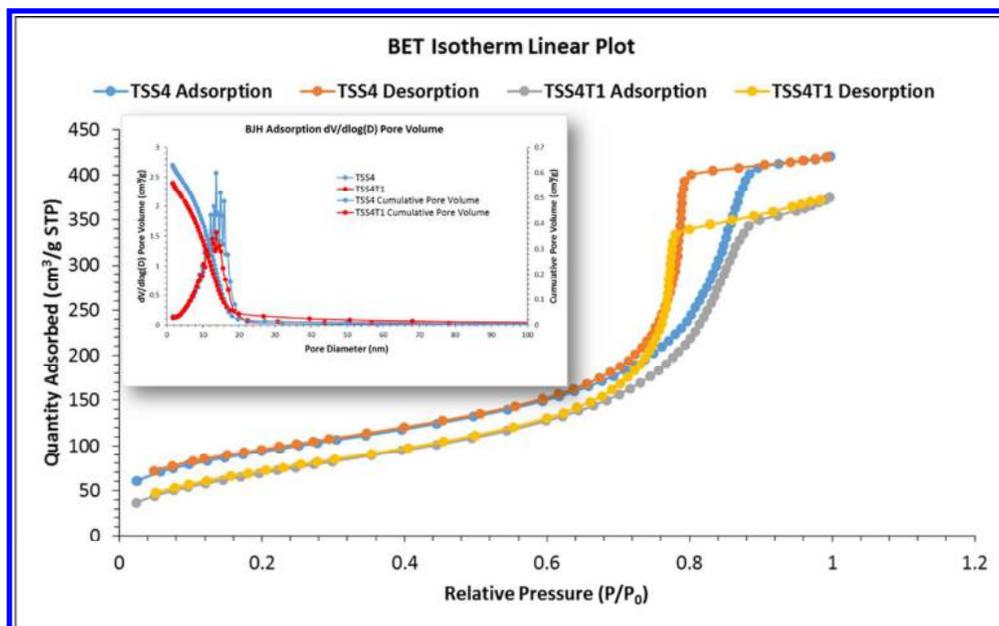


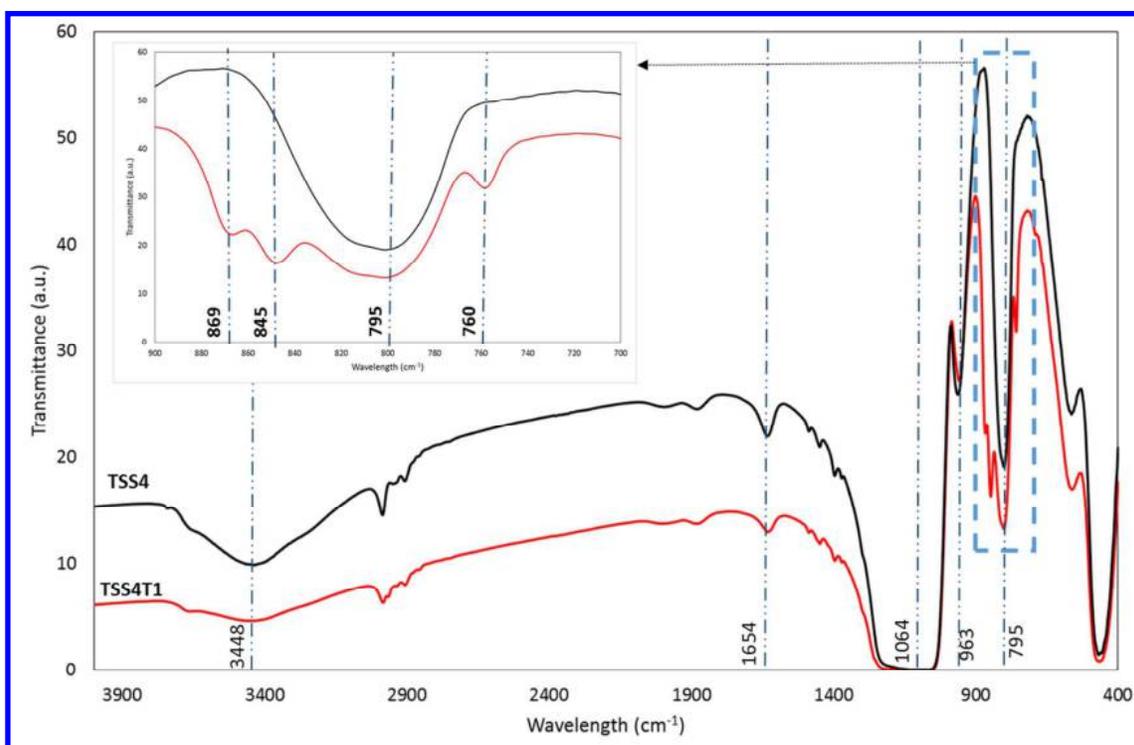
Figure 3: N_2 adsorption-desorption isotherms of silica (TSS4) and functionalised silica (TSS4T1). The insert is the BJH pore distribution of both samples.

Table2: Surface structure parameters of TSS4 and TSS4T1

Sample	Surface Area (m^2/g)	Pore Volume (cm^3/g)	C value	Average Pore Diameter (nm)
TSS4	330.78 ± 2.37	0.647	143.84	7.82
TSS4T1	266.51 ± 1.26	0.571	37.74	8.56

Fig. 4 shows the FTIR spectra of the silica nanoparticles pre- and post-trimethylsilylation to confirm whether the hydrophobic reagent had been grafted on to the surface of the silica nanoparticles. Both spectra have a very strong absorption band at 1064 cm^{-1} and two medium intensity bands at 963 cm^{-1} and 795 cm^{-1} which

are as expected from the silica in the sol-gel process⁴³. The signal at 3448 cm^{-1} is due to free OH group and the signal at 1654 cm^{-1} is due to the antisymmetric vibration of Si-OH. The bending and stretching vibrations of the Si-O-Si bond are represented in the two absorption peaks near 1064 cm^{-1} and 795 cm^{-1} . The absorption peak at 963 cm^{-1} is associated with Si-OH group stretching vibrations. Whilst the FTIR spectra of TSS4 and TSS4T1 exhibit many similarities, two additional absorption peaks appear for TSS4T1, representing the Si-CH₃ stretching and bending vibrations at 845 cm^{-1} and 760 cm^{-1} ⁴⁴⁻⁴⁵. The band intensity of the free hydroxyl band and the hydrogen bonded hydroxyl groups are significantly decreased for the TSS4T1 sample compared with the TSS4 sample. The spectra shown in Fig. 4 demonstrate the successful silylation of the surface of the silica particles



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3 Figure 4: FTIR spectra of TSS4 and TSS4T1 between 4000 and 400 cm^{-1} at room
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5 temperature. The insert is the magnified wavelength between 700 cm^{-1} and 900 cm^{-1} .
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10 Water vapour sorption study on TSS4 and TSS4T1

11 Fig. 5A shows the sorption isotherms of water vapour on TSS4 and TSS4T1 samples.

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13 The absolute hysteresis (obtained by subtracting the adsorption from desorption
14 isotherm loop moisture content values) is given in Fig. 5B. The type of water vapour
15 sorption according to the IUPAC classification for TSS4 is a type IV isotherm,
16 whereas it is a type I(b) isotherm for TSS4T1⁴¹. Type I isotherms are seen in
17 microporous solids which have relatively small external surfaces. The classification
18 of Type I isotherms can be categorised further as Type I(a) which represents
19 microporous materials with predominantly narrow micropores less than 1nm wide,
20 and Type I(b) which represents those materials with a larger range of pore sizes from
21 wider micropores to narrow mesopores. Type IV isotherms are typically found in
22 mesoporous adsorbents. In mesopores the adsorption behaviour is driven by the
23 adsorbent-adsorptive interactions and the interactions between molecules when in
24 the condensed state³⁹. In this case, although the TSS4 and TSS4T1 have a similar pore
25 size distribution, the interactions between the adsorbent-adsorptive have been
26 changed due to the presence of TMS group on the surface of TSS4T1. The amount of
27 water vapour adsorbed on the samples at the same relative pressure decreases
28 significantly after surface modification, especially at higher RH levels. There is a
29 significant reduction in hysteresis between TSS4 and TSS4T1 in the upper part of the
30 hygroscopic range. The reduction of water vapour adsorption and hysteresis on the
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3 TSS4T1 sample is attributed to the superficial hydroxyl groups (-OH) being replaced
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5 by silazane (Si-N) groups, which is also evident in the FTIR curves (Fig. 4). The
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7 water vapour isotherm for the TSS4T1 sample still shows a clear knee at lower RH
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9 levels, and therefore does not display the typical type III isotherm characteristics,
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11 which would have been expected where water adsorption occurs on a hydrophobic
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13 surface⁴². This may be because TSS4T1 has many active sites for water vapour
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15 adsorption despite the silylation which is indicative of residual silanols, a finding
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17 which is supported by the FTIR analysis. The absence of capillary condensation of
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19 water is also of note in the TSS4T1 samples, as can be seen in Fig.5, where the
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21 TSS4T1 isotherm shows minimal mass increase and no measurable hysteresis, in
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23 contrast to TSS4. The possible reason is that the increased hydrophobicity decreases
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25 surface energy of pore walls and thus decreases the tendency for capillary
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27 condensation.
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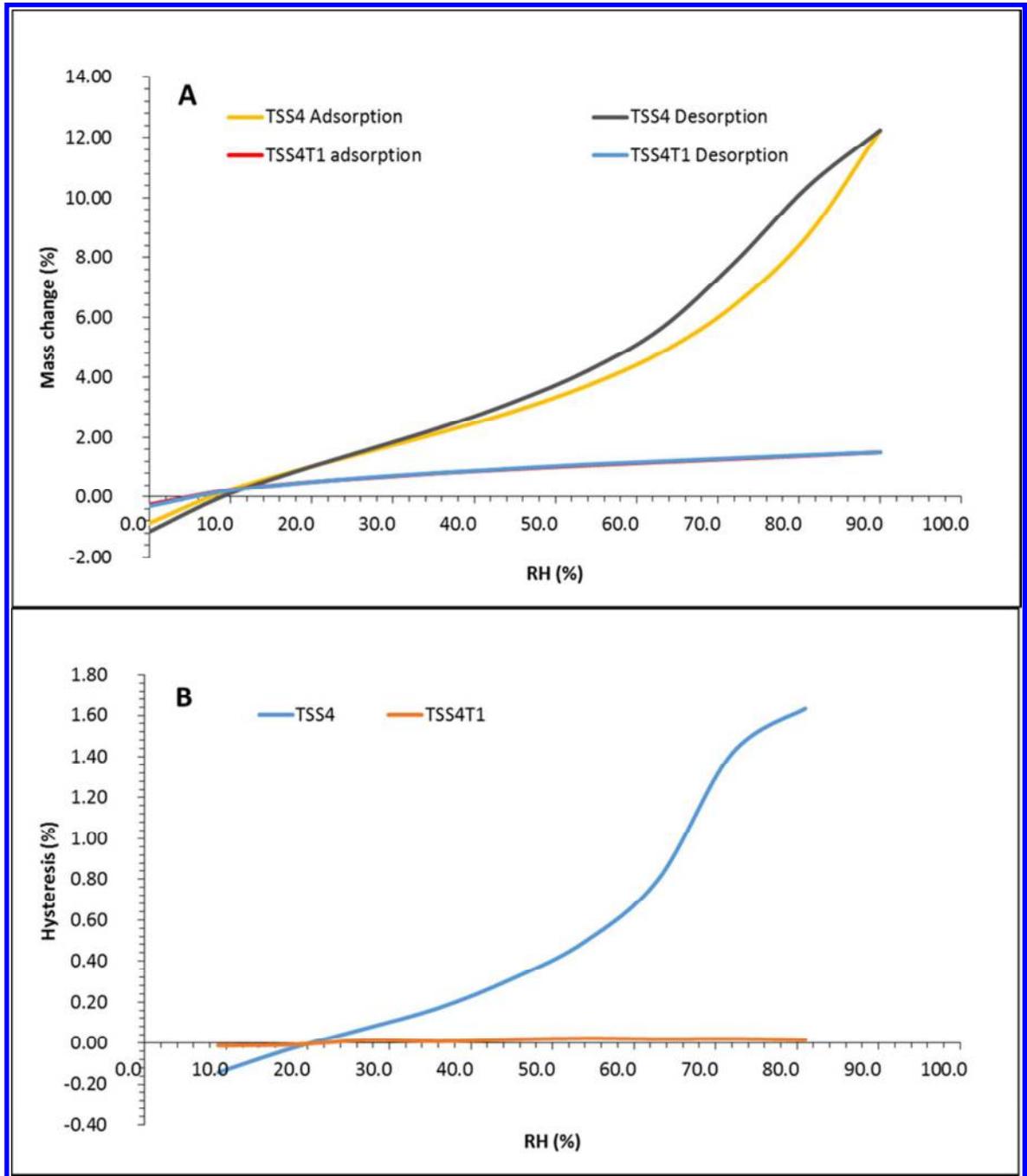


Figure 5: *Water vapour adsorption-desorption isotherms of TSS4 and TSS4T1 (A) and absolute hysteresis plotted as a function of RH.*

Characterisation of the hemp shiv and coated hemp shiv

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3 The surface morphologies of the hemp shiv samples were observed by FESEM. Fig. 6
4 shows characteristic views of the raw hemp shiv and the dip-coated hemp shiv. The
5 raw hemp shiv shown in Fig. 6, A and B, exhibits a microscale cell structure (pits)
6 with a typically smooth surface. The diameter of pits is around 2 μm . Fig. 6, C and D
7 show even coating layers of TSS4T1, together with some agglomerated nanoparticles.
8 The coating is cracked at the edges of the pits. The coating layers masked the
9 anatomical details, and the pits are in some cases filled in with TSS4T1 but not
10 completely occluded by the particles. The surface of the coating is very rough
11 according to Fig. 6, E and F. These particles arise from the assembly and aggregation
12 of the silica nanoparticles on the surface of the hemp shiv. The particle size of
13 modified silica were around 20 nm according to the TEM data (Fig.1). There are a
14 great number of voids among these silica nanoparticles. The voids serve to create the
15 conditions required for a hydrophobic coating, as they are able to repel water and
16 trap air.
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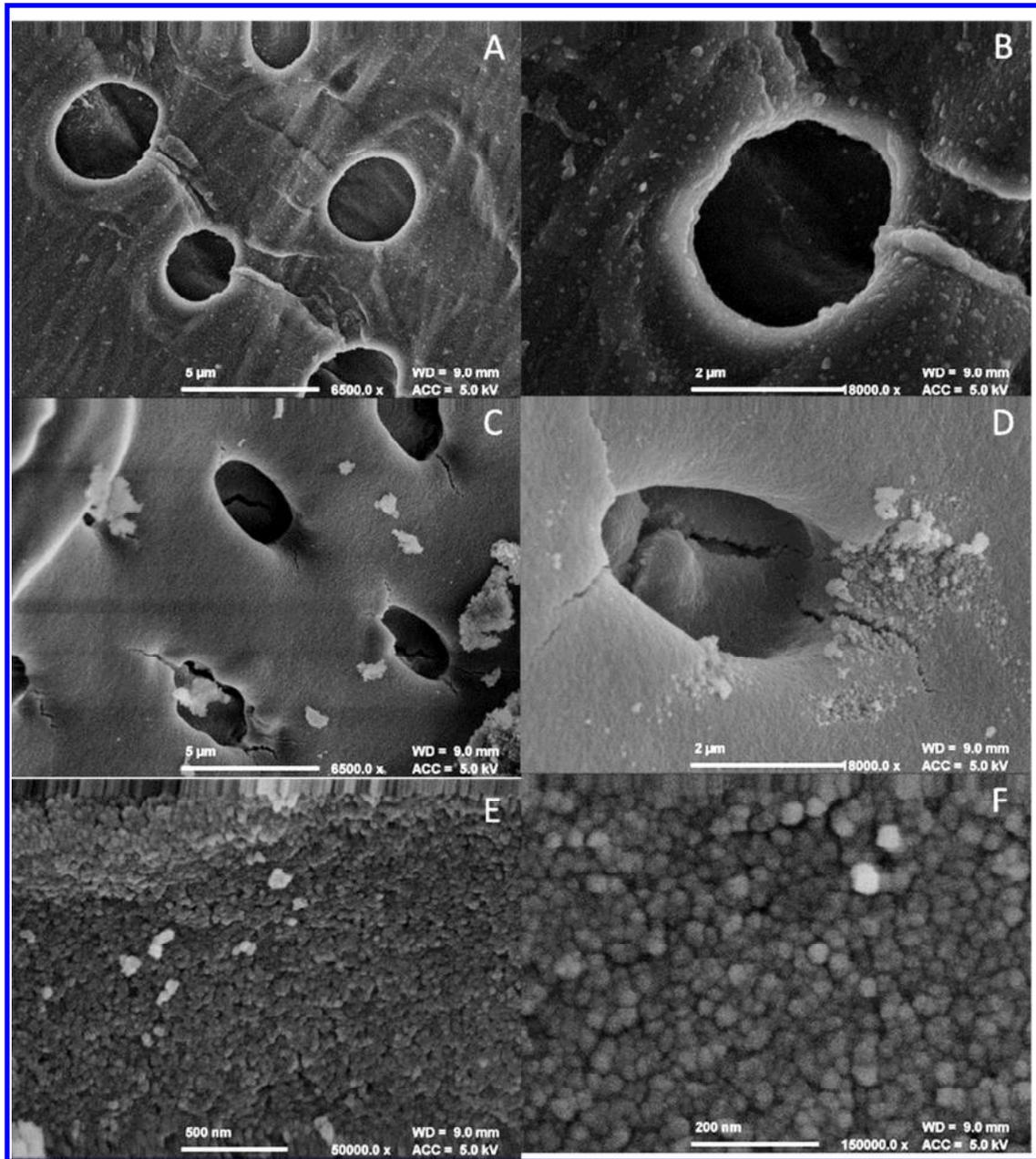
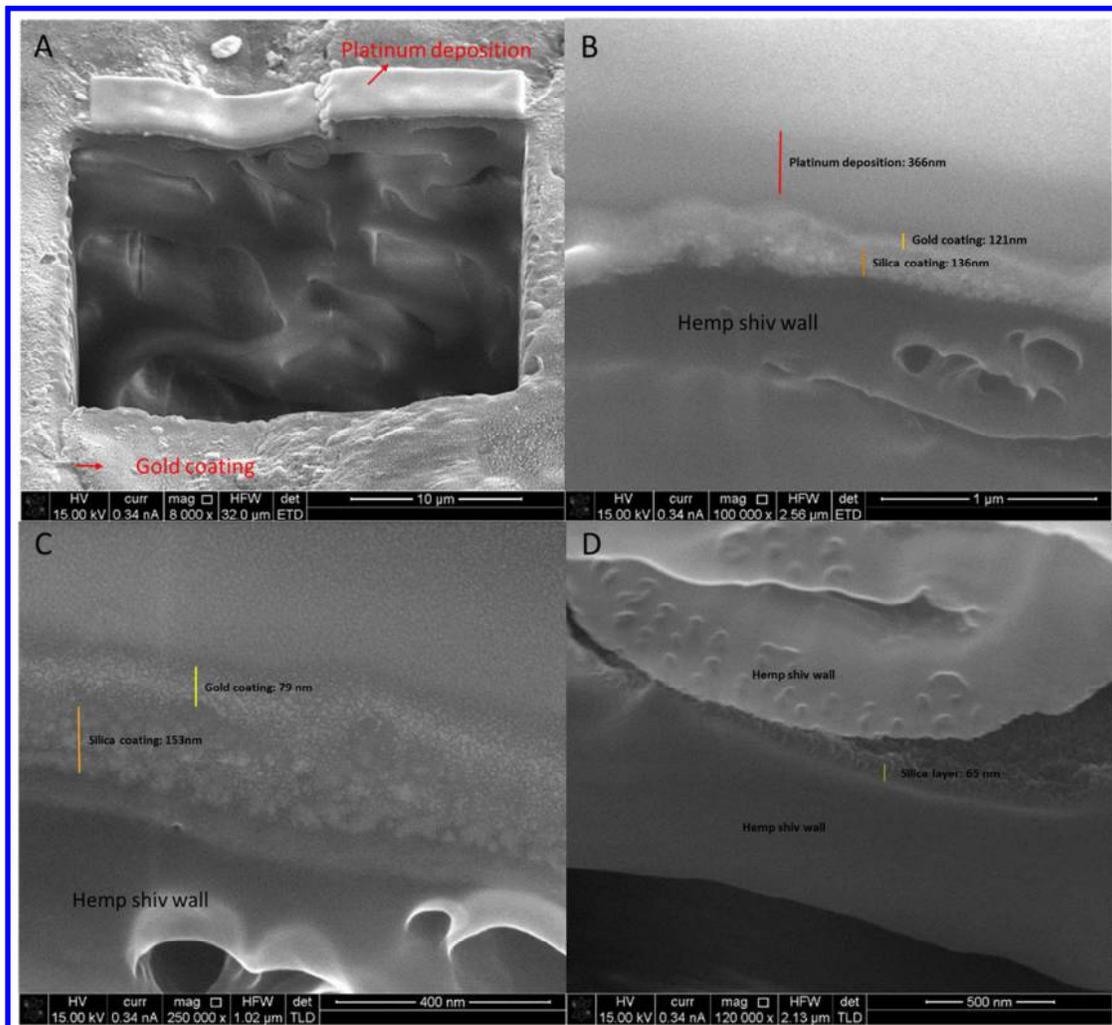


Figure 6: Field emission SEM micrographs of (A) and (B) untreated hemp shiv (C), (D), (E) and (F) dip-coated hemp shiv

Fig.7 shows 'slice and view' on dip-coated hemp shiv using the FIB-SEM. Fig. 7A images a 25 μm diameter hole cut using an ion beam. Fig. 7B and 7C reveal the hidden layered structure of the coating layer. The modified silica coated on the

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3 surface of hemp shiv with multiple layers has a thickness between 130 and 160 nm.
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5 A sputtered gold coating was also applied to increase electrical conductivity for
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7 imaging and platinum deposition for protection of samples is also visible. Fig. 7D
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9 shows that the cell lumens were empty, and a thin layer of silica coating with a
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11 thickness of 65 nm was observed on the surface of the void cell wall. It indicates that
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13 the modified silica not only coated the exterior of the hemp shiv but also penetrated
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15 into the hemp shiv and was deposited onto the internal cell walls.
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3 Figure 7: FIBSEM cross-section micrographs of dip-coated hemp shiv. (A) the surface of dip-
4 coated hemp shiv with gold coating and platinum deposition as viewed using a SEM. The
5 dark region in the centre of the image represents a microscopic hole where material has been
6 machined out by a beam of ions. (B) and (C) magnified views of the previous image reveals
7 the hidden layered structure of a coating layer and clear silica and gold nano-films. (D) voids
8 in the hemp shiv are more clearly shown in this cross-sectional image. A very thin coating
9 layer was observed just above the hemp shiv cell wall.
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21 Fig. 8 shows typical FTIR spectra in an ATR mode for the control and coated hemp
22 shiv samples. The chemical composition of hemp shiv is mainly cellulose,
23 hemicellulose and lignin. Fig. 8 shows a raw hemp shiv FTIR spectrum with broad
24 bands at 3448 cm^{-1} (-OH stretch from polysaccharides), 2918 cm^{-1} (-C-H stretch in
25 methyl and methylene groups from polysaccharides), 1733 cm^{-1} (C=O stretch in
26 unconjugated ketone, carbonyl and ester group from hemicellulose), 1605 cm^{-1}
27 (Aromatic vibration, C=O stretch), 1417 cm^{-1} (CH_2 bending, C=C stretching in
28 aromatic group from lignin), 1369 cm^{-1} (CH_2 bonding from polysaccharides), 1319
29 cm^{-1} (C-H vibration from cellulose) and 1241 cm^{-1} (C= stretch from lignin and xylan).
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The fingerprint region is dominated by bands at approximately $1050\text{ -}1030\text{ cm}^{-1}$ attributed to C-O deformation from various polysaccharide vibrations⁴⁶⁻⁴⁸. The characteristic peaks of the dip-coated hemp shiv are shown in Fig. 8. It can be seen that the absorption peaks at $795, 845, 963$ and 1064 cm^{-1} , are in good agreement with the observed absorption peaks from the pure coating material analysed in Fig. 4. As reported in the literature, the peak at 1064 cm^{-1} represented the asymmetric stretching of the oxygen atoms of two adjacent Si-O-Si units. The bending vibrational

mode of Si-O-Si units is seen in the broad peak at 795 cm^{-1} and the presence of Si-OH group stretching vibrations create the peak at 963 cm^{-1} ⁴⁹⁻⁵¹. The band at 3448 cm^{-1} (-OH stretch), seen on the raw shiv, has almost disappeared on the coated shiv. This is attributed to the replacement of the surface OH groups with trimethyl [-(CH₃)₃] groups²². The FTIR data confirms that the functionalised silica NPs coated on hemp shiv significantly alter the surface chemistry.

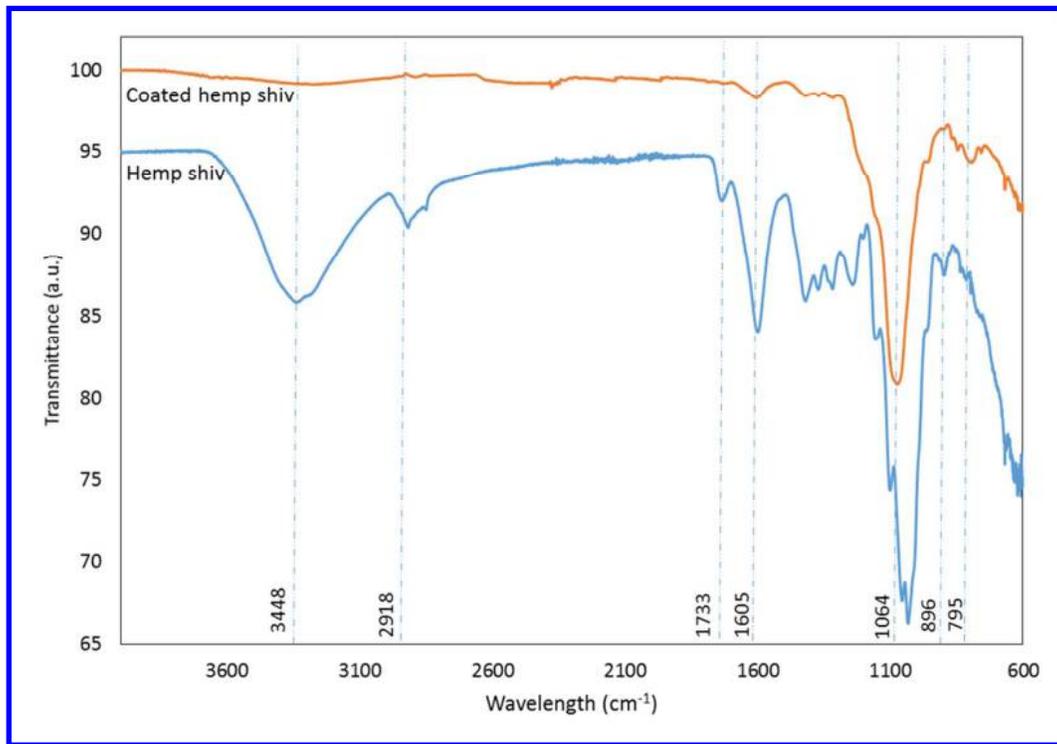


Figure 8: FTIR spectra of the control hemp shiv and the dip-coated hemp shiv

The composition of the untreated and treated hemp shiv was determined by energy dispersive spectroscopy (EDX) analysis. Fig 9A clearly shows the main peaks detected for hemp shiv were carbon and oxygen in the EDX spectrum plus Au peak, from the sputtered coating for SEM preparation, dip-coated hemp shiv also shows a

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3 high silicon content attributable to the TSS4T1 as shown in Fib.9B. By considering
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5 the FTIR data in conjunction with the SEM/EDX data, it appears that the modified
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7 silica coating is linked to the cell wall. In order to confirm this, further studies
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10 utilising solid state nuclear magnetic resonance spectroscopy are required on the
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12 link between the silica nanoparticles and cell wall components
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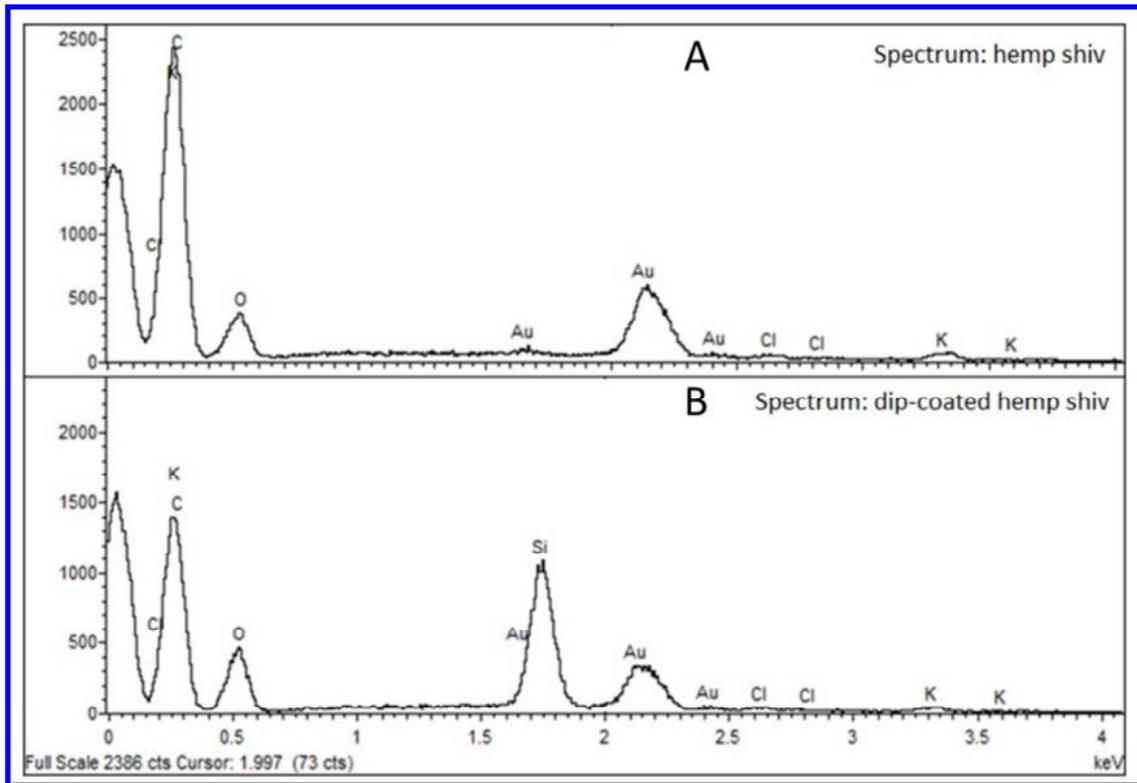
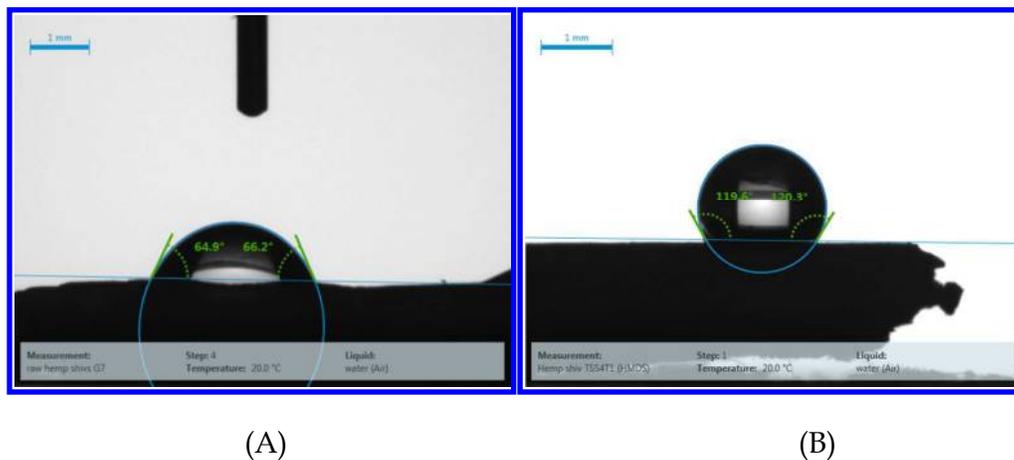


Figure 9: EDX spectra of (A) untreated hemp shiv and (B) dip-coated hemp shiv

Water contact angle and water sorption

In order to identify a change in the surface property of hemp shiv after TSS4T1 coating, the water contact angles of hemp shiv and coated hemp shiv were measured. Fig. 10A shows the hemp shiv has a hydrophilic property with a water contact angle of approximately $65.0^\circ \pm 0.9^\circ$. When the surface of hemp shiv was changed to be

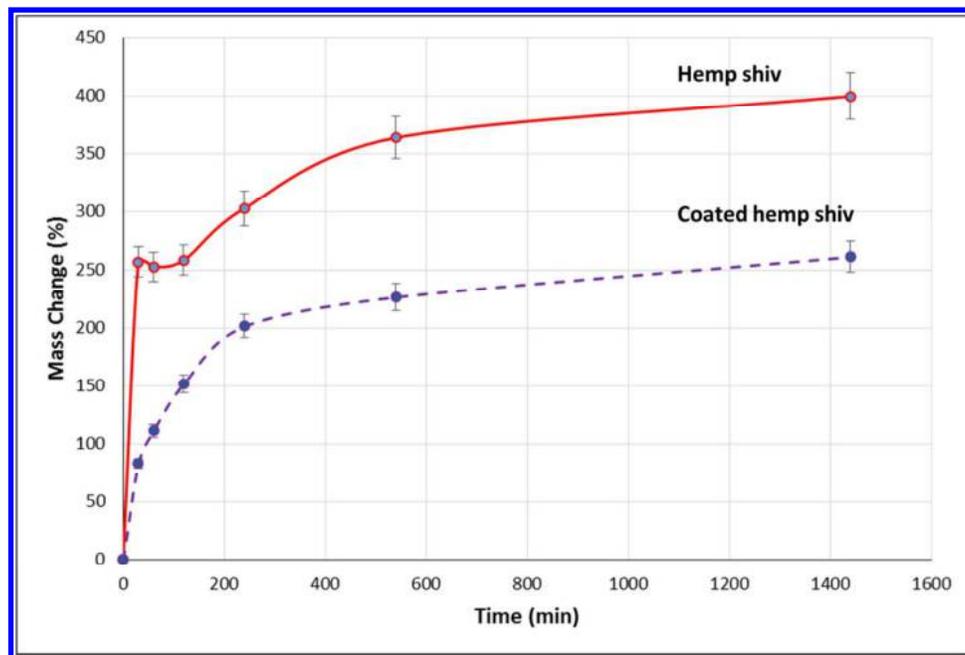
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3 hydrophobic as a result of treatment with TSS4T1, the water contact angle increased
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5 to $120.0^\circ \pm 1.2^\circ$, which is much higher than that of the raw hemp shiv. The readings
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7 reported were the average of five measurements. The water contact angle of coated
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9 hemp shiv was re-measured after immersing the shiv in DI water for 24 hours with
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11 constant stirring in the beaker. The resulting water contact angle of $115.0^\circ \pm 0.8^\circ$ as
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13 shown in the supporting document, is only marginally lower than that achieved in
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15 the first set of measurements, confirming the robustness of the coating. The
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17 equipment used has a reproducibility of $\pm 2^\circ$.
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39 Figure 10: *Water contact angle of raw hemp shiv (A) and coated hemp shiv (B). Note: Image*
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41 *is one of five measurements*
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46 The hydrophobic silica coating acts as a barrier to liquid water to reduce the water
47 sorption. Fig. 11 shows the rate of distilled water absorption for raw hemp shiv and
48 hemp shiv coated with TSS4T1 when held at room temperature over a period of 24 h.
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50 The raw hemp shiv had a maximum water absorption of approximately 400%
51 increase in mass, and absorbed more than half of this amount within the first few
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3 minutes of dipping. The highly porous structure and hydrophilic characteristics of
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5 hemp shiv are responsible for the water absorption performance. A functionalised
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7 silica coating layer on the hemp shiv reduced the water absorption from 400% for the
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9 untreated shiv to 250% for the treated shiv, which is 150% less than the uncoated
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11 shiv. This reduction in water absorption is due to a large decrease in the absorption
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13 of water within the cell walls, with the remaining water being held within the
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15 capillaries of the highly open structure. This demonstrates that treating the shiv with
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17 functionalised hydrophobic silica provides enhanced water resistance and hence
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19 resistance to bio-decay is assumed.
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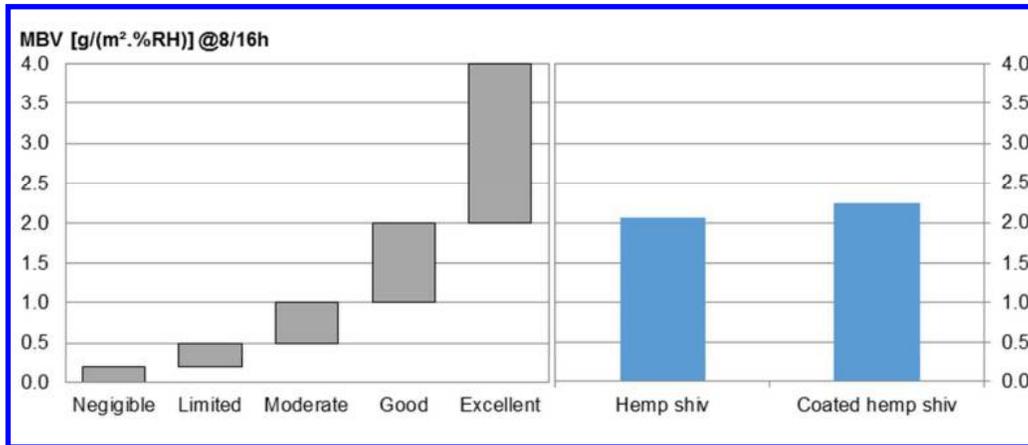


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47 Figure 11. *Dependence of water absorption on time for hemp shiv and coated hemp shiv*

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52 Moisture sorption analysis
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3 Moisture buffering is an energy-efficient material characteristic which enables the
4 absorption, storage and release of moisture, which in turn can regulate the indoor
5 environment by avoiding extremes of humidity and reducing excessive variations. It
6 is a very important parameter for building materials. The MBV values for the hemp
7 shiv and coated hemp shiv are shown in Fig. 12. The coated hemp shiv shows a MBV
8 value of 2.25 g/(m².%RH), while the MBV value of hemp shiv is around 2.07
9 g/(m².%RH). A slight increase of MBV is obtained with the coating. Therefore,
10 according to the Nordtest classification, both untreated and treated hemp shiv
11 exhibit excellent hydric regulation properties, which remain in the level of moisture
12 buffering (MBV > 2). The research undertaken by Lozhechnikova et al. has shown
13 that the use of wax particles can preserved moisture buffering ability, and may even
14 deliver improvements, describing an example where the MBV was increased to 1.08
15 and 1.09 on the radial and tangential surfaces respectively, whereas solid wax films
16 are seen to reduce the moisture buffering ability to limited levels (0.2 < MBV < 0.5)
17 or even less³⁰. In addition, Lozhechnikonva et al. also reported CA values around
18 160 degrees and MBV around 1.4 using wax particles with layer-by-layer
19 deposition³¹. From these data, It can be seen that particles of functionalised silica or
20 wax offer substantial improvements in the preservation of the efficiency of wood as
21 a natural moisture buffering material when compared to the performances offered
22 by a continuous wax layer. Similarly, commercial treatments such as linseed oil
23 coating and lacquer coating were also found to decrease moisture buffering to
24 moderate or limited levels. The hydrophobic coatings form an impermeable barrier
25 on the hemp shiv that water droplets cannot penetrate. The noticeable advantage of
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3 the functionalised silica NPs or wax particles is that they are too small to block the
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5 primary pores found in the hemp shiv, therefore water vapour can permeate freely,
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7 thus allowing the excellent moisture buffering properties of the underlying material
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9 to be retained.
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29 Figure 12: *Practical moisture buffer value classes and moisture buffer values of hemp shiv*
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31 *and coated hemp shiv*
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36 Conclusions

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38 In conclusion, this work has shown that a simple dip-coating method can create a
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40 coating of functionalised silica nanoparticles which enables the fabrication of
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42 hydrophobic hemp shiv. Uniform and multiple layers of silica coating were achieved
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44 on the surface of both external and internal voids of hemp shiv by using a dip-
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46 coating method. FTIR analysis showed that trimethylsilyl groups were introduced
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48 onto the surface of silica nanoparticles and replaced the hydroxyl groups (-OH). The
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50 specific surface areas of silica NPs decreased after modification by HMDS. The
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52 surface modification decreased the moisture adsorption content and reduced the
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3 level of hysteresis shown between the isotherms for adsorption and desorption. In
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5 addition, the results revealed that the functionalised silica nanoparticles were
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7 deposited uniformly over the hemp shiv surface with multiple layers. The initially
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9 hydrophilic hemp surface was made hydrophobic with the silica treatment, and
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11 exhibited a water contact angle of 120° or higher. The functionalised silica coating
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13 deposited on the hemp shiv surface led to a reduction in water absorption of 150% of
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15 the mass of the hemp shiv (from 400% for the untreated shiv to 250% for the treated
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17 shiv) whilst still retaining excellent moisture buffering properties of the hemp shiv.
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24 **Competing interest**

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26 We declare we have no competing interests.
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31
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36 programme, under grant agreement No. 636835.
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42 **Supporting Information**

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44 Durability test of silica-coating on the hemp shiv
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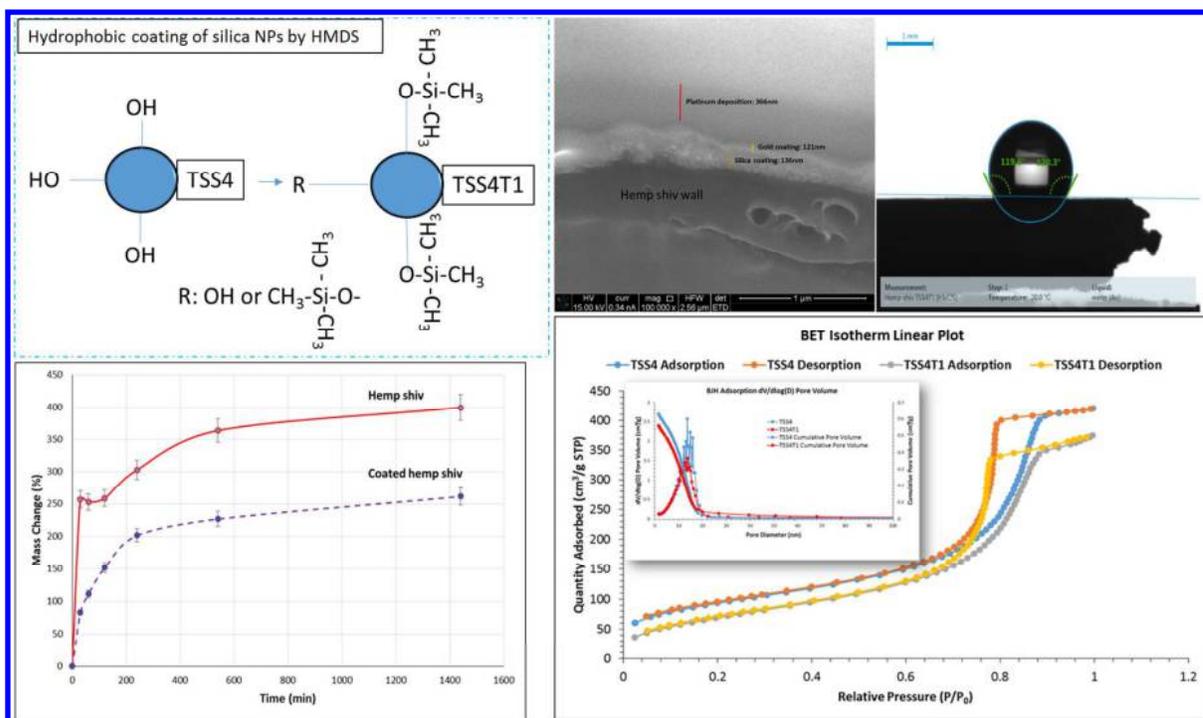
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Synopsis:

This work develops bio-based materials for construction to offer a number of benefits as well as desirable environmental characteristics such as sustainability and renewability.