



HAL
open science

Predicting experimentally stable allotropes: Instability of penta-graphene

Christopher P. Ewels, Xavier Rocquefelte, Harold W. Kroto, Mark J. Rayson, Patrick R. Briddon, Malcolm I Heggie

► **To cite this version:**

Christopher P. Ewels, Xavier Rocquefelte, Harold W. Kroto, Mark J. Rayson, Patrick R. Briddon, et al.. Predicting experimentally stable allotropes: Instability of penta-graphene. Proceedings of the National Academy of Sciences of the United States of America, 2015, 112 (51), pp.15609-15612. 10.1073/pnas.1520402112 . hal-01240650

HAL Id: hal-01240650

<https://univ-rennes.hal.science/hal-01240650>

Submitted on 7 Jan 2016

HAL is a multi-disciplinary open access archive for the deposit and dissemination of scientific research documents, whether they are published or not. The documents may come from teaching and research institutions in France or abroad, or from public or private research centers.

L'archive ouverte pluridisciplinaire **HAL**, est destinée au dépôt et à la diffusion de documents scientifiques de niveau recherche, publiés ou non, émanant des établissements d'enseignement et de recherche français ou étrangers, des laboratoires publics ou privés.

Predicting experimentally stable allotropes: Instability of Penta-graphene

Christopher P. Ewels^{*1}, Xavier Rocquefelte⁺, Harold W. Kroto⁺¹, Mark J. Rayson^{‡*}, Patrick R. Briddon^{*}, Malcolm I. Heggie[‡]

^{*} Institut des Matériaux Jean Rouxel, CNRS UMR 6502, Université de Nantes, BP32229 Nantes, France

⁺ Institut des Sciences Chimiques de Rennes, UMR 6226 CNRS, Université de Rennes 1, Rennes, France

[†] Department of Chemistry and Biochemistry, 95 Chieftain Way, Florida State University, Tallahassee, Florida 32306

[‡] Department of Chemistry, University of Surrey, Guildford, Surrey GU2 7XH, UK

^{*} School of Electrical and Electronic Engineering, University of Newcastle, Newcastle upon Tyne, NE 1 7RU, UK

This version has been accepted for publication and undergone full peer review but has not been through the copy editing, typesetting, pagination and proofreading process, which may lead to differences between this version and the Version of Record (see DOI [10.1073/pnas.1520402112](https://doi.org/10.1073/pnas.1520402112))

ABSTRACT

In recent years a plethora of theoretical carbon allotropes have been proposed, none of which have been experimentally isolated. We discuss here criteria that should be met for a new phase to be potentially experimentally viable. We take as examples Haeckelites, 2D networks of sp^2 -carbon containing pentagons and heptagons, and “Penta-graphene”, consisting of a layer of pentagons constructed from a mixture of sp^2 - and sp^3 - coordinated carbon atoms. In 2D-projection appearing as the “Cairo pattern”, pentagraphene is elegant and aesthetically pleasing. However we dispute the author’s claims of its potential stability and experimental relevance.

SIGNIFICANCE STATEMENT

We describe criteria that should be applied when evaluating whether theoretically proposed carbon allotropes may be experimentally isolated. We discuss the importance of energetic isomeric “funnels” centered on a stable allotropic form, the role of defects in catalyzing structural transformations to lower energy isomers, and chemical stability. This is demonstrated with literature examples such as C60 and B80. We apply these criteria to a recently proposed carbon allotrope, Penta-graphene, demonstrating with the aid of density functional calculations that it will not be experimentally attainable. A second example, Haeckelites, are unlikely to be experimentally achievable when neutral, but may be stabilized through significant charge transfer. The principals discussed here are general and can be applied to any theoretically proposed materials.

INTRODUCTION

One of the joys of carbon research is the huge flexibility of carbon bonding (1-4), resulting in many varied allotropes that have already been experimentally identified. Computational modeling opens the floor to predicting many more, and tools such as graph theory (5) and evolutionary algorithms (6) allow systematic exploration of potential bonding networks. New computationally proposed phases are typically identified as metastable via positive phonon modes, and sometimes via molecular dynamics simulations showing lattice coherence at experimental operating temperatures. However there are common criteria beyond these two tests which link those allotropes which have been experimentally isolated.

Firstly they occupy deep potential wells in the surrounding energetic landscape. Additionally the surrounding energy wells are all higher in energy, “funneling” towards the stable structural form. Finally, barriers to subsequent conversion to alternative structures are typically high. Buckminsterfullerene, I_h -C₆₀, is a good example. The disconnectivity graph for C₆₀ connecting the 1,812 isomers with pentagonal and hexagonal faces via branches whose height indicates the transformation barrier has a ‘willow tree pattern’, with a gentle funnel running towards the stable I_h -C₆₀ isomer (7) (**Figure 1a**). The relatively high barriers are accessible during high temperature growth, and alternatively can be catalysed via the presence of impurities or carbon interstitial atoms (8-10).

In contrast, attempts to experimentally isolate higher order boron fullerenes have been largely unsuccessful to date. For the proposed fullerene B_{80} this can be understood since the energy landscape was shown to feature many closely related isomers with similar (and sometimes lower) energies (11). In contrast calculations for B_{40} , for which there are first experimental indications (12), show a single (D_{2d} , 1A1) cage isomer, energetically well separated from alternative isomers (**Figure 1b**). This behavior is consistent with the rules discussed above.

We apply here a similar analysis for experimental viability to other proposed phases, starting with “Penta-Graphene”, a two-dimensional carbon allotrope proposed by Zhang et al (4). The structure can be viewed as a series of out-of-plane distorted ethylene units connected via tetrahedral sp^3 -carbon linkers. The result is a corrugated layer which in projection matches the “Cairo pattern” of distorted pentagons (Figure 2a).

RESULTS AND DISCUSSION

Thermodynamic Stability: Relative Energy

The first test of any new proposed structure is of its thermodynamic stability. Considering Penta-graphene, while real phonon energies (positive eigenvalues from the Hessian matrix) indicate that it is at least a local structural minimum (4), its formation enthalpy shows that it is a very high energy structure. We have performed a number of calculations on pentagraphene and structural derivatives, using DFT/LDA calculations (see Method). Penta-graphene is 0.761 eV per atom less stable than graphene. This is significantly less stable than amorphous carbons (0.16eV) (13) most nanotubes and I_h - C_{60} (0.39eV) (14), and places it in a similar energy range to experimentally unconfirmed isomers such as R3-Carbon (15).

Transforming to Graphene. Secondly, penta-graphene is not the stable centre of a “funnel” of isomeric structures, but instead forms part of an energetic funnel of structures centred on graphene. **Figure 2** shows an example sequence of bond rotations and bond-breakages by which penta-graphene can be transformed directly to graphene. Each step is exothermic, but only if the unit cell vectors are allowed to geometrically relax along with all atoms. This does not appear to have been done in the original paper by Zhang *et al.* The unit cell area increases 19.6% from penta-graphene (25.96\AA^2) to graphene (31.05\AA^2). This is because the sp^3 bonding and corresponding non-planarity in penta-graphene give it a relatively dense 2D-projected basal plane compared to graphene. Constraining the unit cell dimensions to those of penta-graphene therefore energetically disadvantages planar sp^2 restructuring.

The periodic boundary conditions for the calculations necessarily require that the set of rotations at each step, e.g. from (a) to (b) be simultaneous as shown, and this may be argued to be improbable. However, we have also tested the equivalent single isolated bond rotation for a 2×2 repeated supercell of penta-graphene. Rotating just one of the C-C bonds marked in **Figure 2a** releases 0.812eV, showing that transformation from penta-graphene to graphene could occur stepwise.

Thus since Penta-graphene forms part of a continuous energetic funnel towards Graphene, it will not be possible to isolate experimentally on thermodynamic grounds.

The above analysis is consistent with molecular dynamics simulations which have been used to map out carbon energy landscapes in terms of structural crystallinity and density, as a function of applied pressure and temperature (16). Consistent with experiment, these landscapes predict that stable crystalline carbon polytypes contain only one hybridization state of carbon (either sp^2 or sp^3), with mixed phases stable in amorphous configurations. They also show energy funneling towards fully sp^2 or sp^3 phases, depending on system pressure.

Metastable structures may still nonetheless be isolated if their conversion to stable phases is kinetically inhibited due to high interconversion barriers, so we next consider kinetic stability criteria.

Kinetic Stability and the importance of defects

A classical stability test is the use of molecular dynamics (MD) simulations at experimental temperatures. Penta-graphene was shown in this way to maintain its structure for 1ps at 300K (4). However this is not sufficient to demonstrate kinetic stability.

Structural distortion in materials is typically a localized process commencing at defective sites, often catalysed by impurities and defects. The 90° carbon-carbon bond rotation processes shown in Figure 2a are an example of this. Each bond rotation step represents the annihilation of a dislocation dipole in the underlying graphene lattice, and requires the simultaneous breaking and reformation of two carbon bonds. Similar carbon-carbon bond rotation in graphene has an extremely high calculated enthalpy barrier of 8.99eV (9). For the single bond rotation step in penta-graphene discussed above we have calculated the reaction barrier to be in the range 2.33-3.04 eV (17). This is already significantly lower than that of graphene.

However the C-C bond rotation barrier in graphene is reduced by a factor of four in the presence of defects such as carbon adatoms (9), and similarly introducing a carbon adatom to penta-graphene also drops the calculated barrier by 35% to only 1.51 eV. Additionally the local energy release associated with this restructuring will also likely render the process auto-catalytic. Thus in the presence of any defects penta-graphene is unlikely to be stable. We note that the authors did indeed run MD simulations for point defects in Penta-Graphene (4), but these were for only 5ps and with fixed lattice constants as discussed above.

Chemical Stability and Oxidation

While chemical stability is not strictly a criterion for experimental viability, it is nonetheless an important indicator. In order to examine potential environmental stability of Penta-graphene, we calculated its reaction with O₂. In the triplet state O₂ oxidation of the surface is highly exothermic, releasing 2.24eV per O₂ molecule, with a barrier to chemisorption of only 0.16eV. Accounting for spin conversion to an eventual singlet state (surface crossing) releases a further 2.34eV. Oxygen forms epoxides at the distorted localized surface C=C bonds. Thus Penta-graphene would undergo highly exothermic spontaneous oxidation on the slightest exposure to air.

Haeckelites and related sp²-layered structures

Our next example is the Haeckelites (1-3), a family of layered sp²-carbon structures constructed from pentagons, heptagons, and optionally hexagons. Despite extensive theoretical investigation since their first proposition nearly twenty years ago (1), these structures have never been unambiguously experimentally isolated.

Figure 3a-c shows three Haeckelite structures. In terms of thermodynamic stability all three are less stable than graphene, (Fig.3a) Oblique O_{5,6,7} by 0.377eV/C, (Fig.3b) rectangular R_{5,7} by 0.244eV/C and (Fig.3c) Hexagonal H_{5,6,7} by 0.251eV/C, although the energies relative to graphene are significantly lower than that of Penta-graphene. Structures 3a and 3b can both be converted to graphene via 90° bond rotations of the bonds labeled 'X' in Figure 3, and hence lie on the thermodynamic 'funnel' terminating at graphene. As discussed previously, the bond rotation barrier can also be lowered through the presence of defects. Structure 3c cannot reconstruct via bond rotations to Graphene, but if produced through a sequential edge growth process could still be replaced by thermodynamically preferable graphene. Thus from the arguments above, we would not expect these structures to be experimentally viable.

If the system is charged however, the thermodynamic picture changes. Figure 4 shows the energy of Haeckelite-3c and Penta-graphene relative to graphene at different charge states. While Penta-graphene becomes increasingly unstable with charging, the Haeckelites become increasingly stable, and for high charge accumulation, Haeckelite-3c becomes more stable than graphene. The effect is even more marked for the Octahedral-square structure (Figure 3e) which shifts from 0.57eV/C less stable than neutral graphene, to comparable energy when charged 0.25e/C.

Structural analogues that have indeed been experimentally isolated are metal borocarbides MB_2C_2 , where $M=Mg, Sc, Ce, Ca, Y, La, Lu$ (18,19). These are layered B_2C_2 structures intercalated with metal ions (Figure 3). Hexagon, pentagon-heptagon (Figure 3b) and octagon-square lattices (Figure 3d) have been experimentally identified depending on the metal cation (19). Among these phases MgB_2C_2 exhibits a stable hexagonal graphene-like lattice. This can be understood since the Mg^{2+} cation donates sufficient electrons to compensate the electron deficiency of boron with respect to carbon, i.e. 2 electrons for each B_2C_2 motif. In contrast M^{3+} cations ($La^{3+}, Ce^{3+}, Y^{3+}, Sc^{3+}, Lu^{3+}$) donate one extra electron compared to Mg^{2+} , leading to an extra charge of 0.25e/lattice atom. These cations form non-hexagonal borocarbide lattices, in agreement with our analysis of thermal stability vs charge state for carbon sheets, where at 0.25e/carbon atom Haeckelite and octahedron-square lattices become thermodynamically competitive with graphene (20).

Thus while the calculations suggest that isolated Haeckelites are unlikely to be formed experimentally, in situations of heavy doping, such as metal-ion intercalation in layered crystals, or ion overlayers on surfaces, experimental isolation of Haeckelites may be possible.

CONCLUSIONS

The calculations presented here strongly suggest that penta-graphene will not be an experimentally achievable allotrope of carbon. Not only would penta-graphene be difficult to isolate from the plethora of alternative isomers with similar energies, it should rapidly restructure towards graphene in the presence of even a few catalytic impurities. Even were it to form, it would not be environmentally stable. Haeckelites, while unlikely to be experimentally attainable in isolation, may still be experimentally achievable in the presence of charge transfer ions.

In general when determining the experimental feasibility of synthesizing new carbon phases, it is not sufficient to establish the pristine material as a metastable minimum on a local energy surface. We have highlighted here the importance that the structure lie at the apex of a disconnectivity graph of related isomers, occupying a unique energetically isolated position. Chemical and kinetic stability, notably in the presence of catalytic defects (both intrinsic and extrinsic) are also critical. This analysis could easily be extended to other proposed carbon allotropes, such as the various graphyne phases (21), and indeed is general beyond its simple application to carbon.

METHOD

Quantum Chemical Calculations.

DFT calculations were performed using the AIMPRO code (22), with a basis set containing 22 independent Gaussian-based functions per carbon atom and 40 per oxygen atom. Hartwigsen-Goedecker-Hutter relativistic pseudopotentials were used (23). Finite temperature smearing was used for the electronic state population with temperature $kT=0.01eV$. Lattice vectors were relaxed simultaneously with atom positions and lattice symmetry broken. For Pentagraphene and related 12 atom unit cells a 4×4 Monkhorst-Pack (24) k-point grid was used, for the 2×2 48 atom cell a 2×2 k-point grid was used. Saddle points were determined using the climbing nudged elastic band algorithm (25). Orthorhombic octahedron-square cells contained 4 atoms, $8\times 8\times 1$ k-points, hexagonal H_{567} cells had 16 atoms, $4\times 4\times 1$ k-points, $R_{5,7}$ cells contained 48 atoms, $2\times 2\times 1$ k-point grid, and $O_{5,6,7}$ cells 12 atoms, $4\times 4\times 1$ k-point grid.

ACKNOWLEDGMENTS. This project has received funding from the European Union's Horizon 2020 research and innovation programme under the Marie Skłodowska-Curie grant agreement No 642742, National Science Foundation (NSF) through the Division of Materials Research (DMR) and Division of Chemistry (CHE), grants NSF DMR-1157490 and CHE-1019193, the Florida State University Research Foundation, and the Swedish Research Council (Reg. no. 2012-3174).

Author contributions: C.P.E., H.W.K., X.R. M.I.H. designed research; C.P.E., X.R. performed research; M.J.R., P.R.B. developed calculation tools; C.P.E., X.R., H.W.K., M.I.H. wrote the paper.

The authors declare no conflict of interest.

¹ To whom correspondence may be addressed. E-mail: chris.ewels@cnsr-immn.fr, kroto@chem.fsu.edu

REFERENCES

- Crespi VH, Benedict LX, Cohen ML & Louie SG (1996) Prediction of a pure-carbon planar covalent metal. *Phys. Rev. B* 53: R13303.
- Terrones H, Terrones M, Hernandez E, Grobert N, Charlier J-C & Ajayan PM (2000) New metallic allotropes of planar and tubular carbon. *Phys. Rev. Lett.* 84: 1716.
- Rocquefelte X, Rignanese GM, Meunier V, Terrones H, Terrones M & Charlier JC (2004) How to Identify Haeckelite Structures: A Theoretical Study of their Electronic and Vibrational Properties. *Nano Letters* 4 (5):805-810.
- Zhang S, Zhou J, Wang Q, Chen X, Hawazoe Y & Jena P (2015) Pentagraphene: A new carbon allotrope. *Proc. Nat. Acad. Sci.* 112 (8): 2372-2377. doi:10.1073/pnas.1416591112
- Strong RT, Pickard CJ, Milman V, Thimm G & Winkler B (2004) Systematic prediction of crystal structures: An application to sp³-hybridized carbon polymorphs. *Phys. Rev. B* 70: 045101.
- Zhou R & Cheng Zeng X (2012) Polymorphic Phases of sp³-Hybridized Carbon under Cold Compression, *J. Am. Chem. Soc.* 134 (17): 7530-7538
- Wales D, Miller MA & Walsh TR (1998) Archetypal energy landscapes. *Nature* 394, 758.
- Eggen BR, Heggie MI, Jungnickel G, Latham CD, Jones R, Briddon PR (1996) Autocatalysis during fullerene growth. *Science* 272 (5258): 87-90.
- Ewels CP, Heggie MI & Briddon PR (2002) Adatoms and nanoengineering of carbon. *Chem. Phys. Lett.* 351: 178.
- Dunk PW, Kaiser NK, Hendrickson CL, Quinn JP, Ewels CP, Nakanishi Y, Sasaki Y, Shinohara H, Marshall AG & Kroto HW (2012) Closed network growth of fullerenes. *Nature Communications* 3: 856.
- Pochet P, Genovese L, De S, Goedecker S, Caliste D, Alireza Ghasemi S, Bao K & Deutsch T (2011) Low-energy boron fullerenes: Role of disorder and potential synthesis pathways. *Phys. Rev. B* 83: 081403(R).
- Zhai H-J, Zhao Y-F, Li W-L, Chen Q, Bai H, Hu H-S, Piazza ZA, Tian W-J, Lu H-G, Wu Y-B, Mu Y-W, Wei G-F, Liu Z-P, Li J, Li S-D & Wang L-S (2014) Observation of an all-boron fullerene. *Nature Chemistry* 6: 727-731. doi:10.1038/nchem.1999
- Jacob KT & Seetharaman S (1993) Thermodynamic stability of metallurgical coke relative to graphite. *Metall. Mat. Trans B*, 25B: 149-151. doi:10.1007/BF02663188
- Diogo HP, Minas da Piedade ME, John S, Dennis T, Hare JP, Kroto HW, Taylor R & Walton DRM, (1993) Enthalpies of formation of buckminsterfullerene (C₆₀) and of the parent ions C₆₀⁺, C₆₀²⁺, C₆₀³⁺ and C₆₀⁻, *J. Chem. Soc., Faraday Trans.* 9: 3541-3544, doi: 10.1039/FT9938903541
- Jiang X, Arhammar C, Liu P, Zhao J & Ahuja R (2013) The R3-carbon allotrope: a pathway towards glassy carbon under high pressure. *Scientific Reports*, 3: 1877. doi:10.1038/srep01877
- Powles RC, Marks NA, Lau DWM, McCulloch DG & McKenzie DR (2013) An energy landscape for carbon network solids. *Carbon* 63: 416-422.
- Since we are not able to simultaneously relax the unit cell dimensions with the atom positions within the nudged elastic band calculations, the two quoted values are calculated with the unit cell parameter fixed at the start- and end- structure values respectively. Simultaneous unit-cell and position relaxation may lower this barrier further.
- Smith GS, Johnson Q & Nordine PC (1965) The Crystal Structure of ScB₂C₂, *Acta Cryst.* 19: 668-673.
- Rocquefelte X, Boufelfel SE, Yahia MB, Bauer J, Saillard JY & Halet JF (2005) Structural Preference versus Metal within the MB₂C₂ (M=Mg, Sc, Ca, Y, Ln) Phases : The Coloring Problem Revisited by DFT Calculations, *Angew. Chem. Int. Ed.* 44:7542-7545.
- The exception is Ca²⁺ which also favours a square-octahedral lattice, however this is because of additional effects (ionic radius).
- Malko D, Neiss C, Viñes F & Görling A (2012) Competition for Graphene: Graphynes with direction-dependent Dirac cones. *Phys Rev. Lett.* 108: 086804.
- Rayson MJ & Briddon PR (2009) Highly efficient method for Kohn-Sham density functional calculations of 500-10000 atom systems. *Phys. Rev. B* 80: 205104.
- Hartwigsen C, Goedecker S & Hutter J (1998) Relativistic separable dual-space Gaussian pseudopotentials from H to Rn. *Phys. Rev. B*, 58: 3641.
- Monkhorst HJ & Pack JD (1976). Special points for Brillouin-zone integrations. *Phys. Rev. B* 13: 5188.
- Henkelman G, Uberuaga BP & Jonsson H (2000) A Climbing-Image NEB Method for Finding Saddle Points and Minimum Energy Paths. *J. Chem. Phys.* 113: 9901.

FIGURE LEGENDS

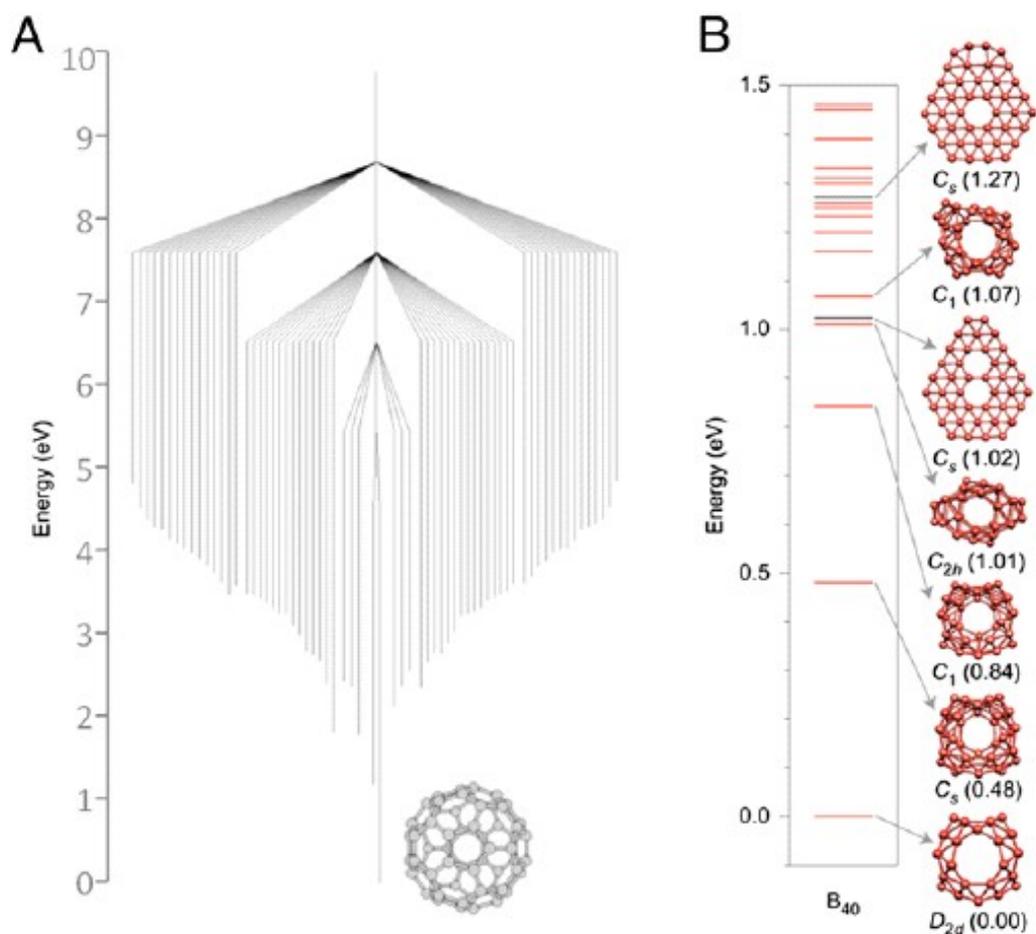


Figure 1. (a) “Willow Tree” pattern of different C_{60} isomers, with the lower points of each vertical bar representing the calculated formation enthalpy relative to I_h-C_{60} , bar heights representing the calculated barrier to transformation. This shows that I_h-C_{60} is significantly more stable than other isomers and lies at the centre of an “energetic funnel”. Adapted from (7). (b) Calculated formation enthalpies of B_{40} , showing the D_{2d} cage structure is significantly more stable than alternative isomers. Taken from Reference (12).

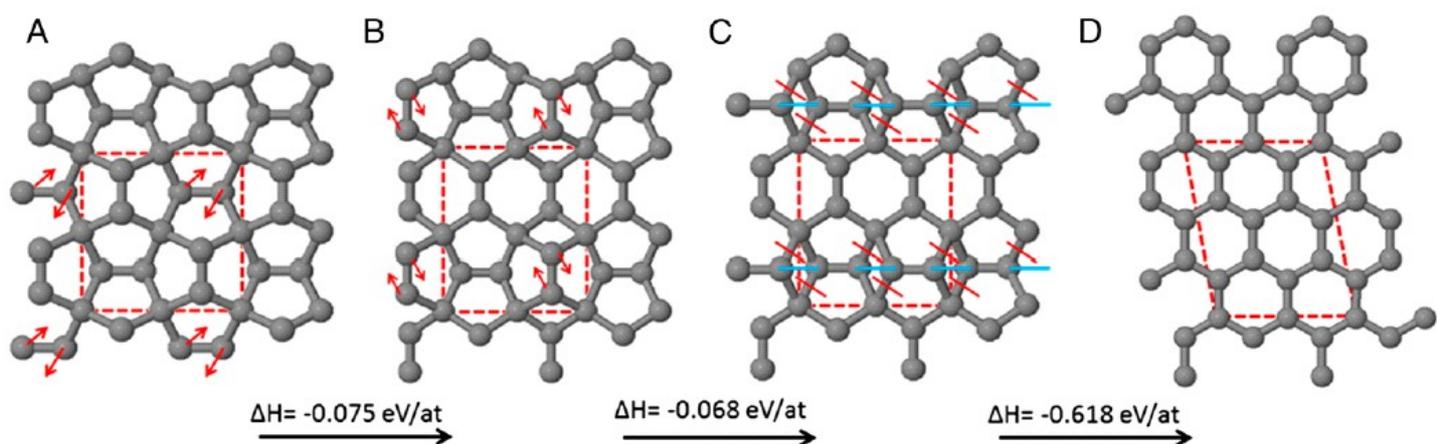


Figure 2. A calculated structural transformation route from (a) penta-graphene to (d) graphene, each step is exothermic. Red arrows indicate direction of motion of atoms for 90° rotation of carbon-carbon bonds. Red (blue) lines indicate C-C bonds that are broken (formed). Note that structures (a-c) were constrained within orthogonal unit cells, this constraint was lifted for step (c) to (d). The final structure, graphene, is 0.761 eV per atom more stable than (a). Unit cells marked with dotted lines, calculated cell dimensions are (a) $5.095\text{\AA} \times 5.095\text{\AA}$, (b) $4.769\text{\AA} \times 5.510\text{\AA}$, (c) $4.888\text{\AA} \times 5.318\text{\AA}$, and (d) $4.883\text{\AA} \times 6.476\text{\AA}$, $\alpha = 100.88^\circ$.

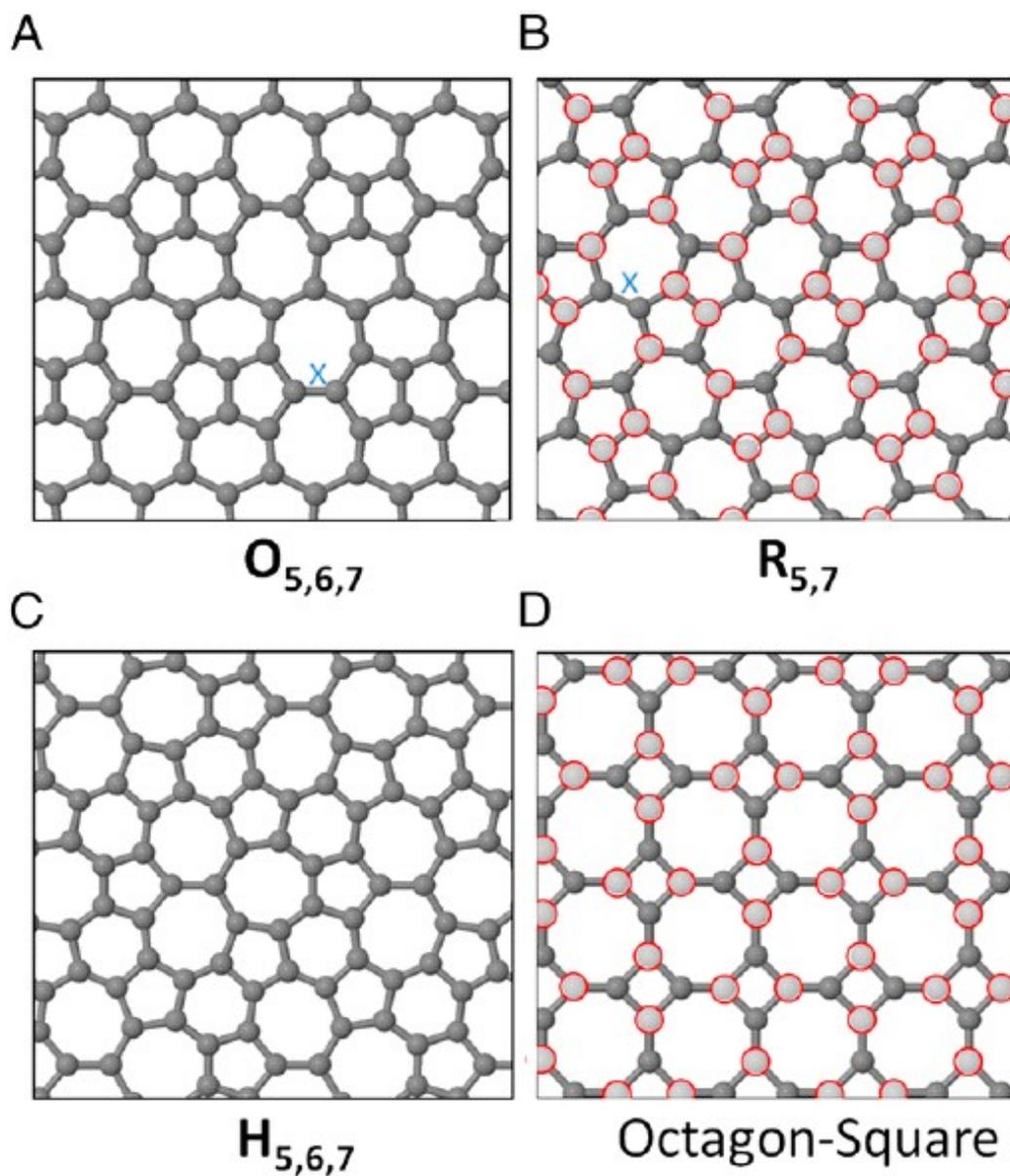


Figure 3. Three Haeckelite structures (a) Oblique $O_{5,6,7}$, (b) a rectangular $R_{5,7}$ and (c) Hexagonal $H_{5,6,7}$ (nomenclature from Ref [1b]), and (d) octagon-square structure. In MC_2B_2 borocarbides, lighter carbon atoms marked with red circles are replaced with boron in (b) $M=Sc$ and (d) $M= Ce, Y, Ca, Ln$ (metal ions located above the centre of heptagons and octagons respectively). Bonds marked 'x', when rotated, convert a 5-7-7-5 patch into four hexagons.

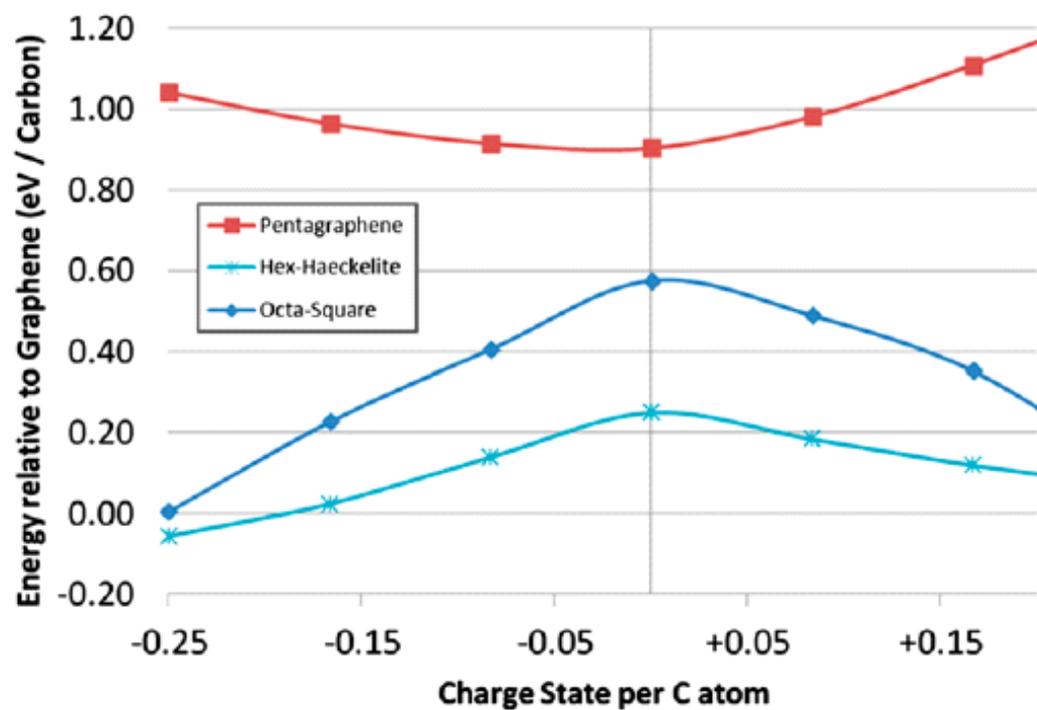


Figure 4. Relative stability of Penta-Graphene (Figure 2a) and Hex-Haeckelite (Figure 3a) compared to graphene, as a function of charge state per C atom. While Penta-Graphene is increasingly unstable with charging, Haeckelites and Square-Octahedral structures are increasingly stabilized, becoming comparable with graphene or even thermodynamically favoured at 0.25e/carbon.